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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17.3.2.1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $-\log 0.100 = 1.000$ ph $-\log 0.100 = 1.000$ b titrant volume 12.50 ml

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web expert answer 100% 2 ratings the example is based upon henderson hasselbalch equation the equation gives relation between ph and pK_a of weak acid the equation is $ph = pK_a + \log \frac{[salt]}{[acid]}$ at equivalence point all a view the full answer transcribed image text

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Volume of NaOH added (mL)	pH
4	6.4
6	6.4
6	5.9
5	9.5
9	5.9

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