



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionacidbasestrong acidweak acid	<ul style="list-style-type: none">concentrationhydrogen ionacidbasestrong acidweak acid
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Figures

- Figure 1: pH scale
- Figure 2: Ion concentration

Points of Consideration

What are the components of acids and bases? How do acids and bases affect the pH of a solution? How do acids and bases affect the solubility of other substances?

What is the difference between a strong acid and a weak acid? How do strong acids and weak acids affect the pH of a solution?

What is the difference between a strong base and a weak base? How do strong bases and weak bases affect the pH of a solution?

What is the difference between an acid雨和 a base雨? How do acids雨和 bases雨 affect the pH of a solution?

Key Equations

$\text{pH} = -\log_{10}[\text{H}^+]$	$\text{pH} = \log_{10}[\text{OH}^-]$
$\text{pH} = 7$	$\text{pH} < 7$
$\text{pH} = 14$	$\text{pH} > 7$
$\text{pH} = 10$	$\text{pH} = 12$
$\text{pH} = 8$	$\text{pH} = 9$
$\text{pH} = 6$	$\text{pH} = 7$
$\text{pH} = 5$	$\text{pH} = 4$
$\text{pH} = 3$	$\text{pH} = 2$
$\text{pH} = 1$	$\text{pH} = 0$

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Eileen Ramsden

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