

Chemistry 2nd Semester Final Exam Review

Chemical Bonds

- Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.
A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.
- Use the information from the previous question to describe the shape and volume for each phase (state) of matter.
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.
 Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.
 Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
- List the three phases of matter in order of increasing intermolecular attractions.
A: Gas, liquid, solid.
- Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?
A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.

Conservation of Matter and Stoichiometry

- The atomic mass unit and the mole are defined using what isotope? *A: Carbon-12*
- Balance the following equations:
 - $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$
A: $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$
 - $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
A: $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 - $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
A: $4\text{P} + 5\text{O}_2 \rightarrow 2\text{P}_2\text{O}_5$
- What isotope is used to define one mole? *A: Carbon-12*
- How many particles (atoms or molecules) are in one mole? *A: 6.02×10^{23}*
- Find the molar mass for the following chemical formulas:
 - ZnF_2 *A: 103.377 g/mol*
 - $\text{Al}_2(\text{SO}_4)_3$ *A: 342.151 g/mol*
 - NH_4IO_3 *A: 192.941 g/mol*
- How many moles are in 4.37 kg of NaOH? *A: 109.3 mol*
- How many atoms are in 2075 g of He? *A: 3.12×10^{26} atoms*
- What volume will 92.31 g of CO_2 occupy at STP? *A: 46.985 L*
- How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?
 $\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$
A: 1.68 g NaOH
- In the reaction, $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$, if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?
A: 126 g MgO

Gases and Their Properties

- What is the kinetic molecular theory (KMT)?
A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
 - Gases are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.*
 - These particles move in a straight line until they collide with another particle or the walls of the container.*
 - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.*
 - There is no force of attraction between gas particles or between the particles and the walls of the container.*
 - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.*

Nd Semester Final Review Questions

Clemens Wendtner



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Table of Contents Nd Semester Final Review Questions

1. Understanding the eBook Nd Semester Final Review Questions
 - The Rise of Digital Reading Nd Semester Final Review Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Nd Semester Final Review Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nd Semester Final Review Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nd Semester Final Review Questions
 - Personalized Recommendations
 - Nd Semester Final Review Questions User Reviews and Ratings
 - Nd Semester Final Review Questions and Bestseller Lists

5. Accessing Nd Semester Final Review Questions Free and Paid eBooks
 - Nd Semester Final Review Questions Public Domain eBooks
 - Nd Semester Final Review Questions eBook Subscription Services
 - Nd Semester Final Review Questions Budget-Friendly Options
6. Navigating Nd Semester Final Review Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Nd Semester Final Review Questions Compatibility with Devices
 - Nd Semester Final Review Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nd Semester Final Review Questions
 - Highlighting and Note-Taking Nd Semester Final Review Questions
 - Interactive Elements Nd Semester Final Review Questions
8. Staying Engaged with Nd Semester Final Review Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nd Semester Final Review Questions
9. Balancing eBooks and Physical Books Nd Semester Final Review Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nd Semester Final Review Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nd Semester Final Review Questions
 - Setting Reading Goals Nd Semester Final Review Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nd Semester Final Review Questions
 - Fact-Checking eBook Content of Nd Semester Final Review Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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