

Chemistry 2nd Semester Final Exam Review

Chemical Bonds

- Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.
A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.
- Use the information from the previous question to describe the shape and volume for each phase (state) of matter.
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.
Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
- List the three phases of matter in order of increasing intermolecular attractions.
A: Gas, liquid, solid.
- Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?
A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.

Conservation of Matter and Stoichiometry

- The atomic mass unit and the mole are defined using what isotope?
A: Carbon-12
- Balance the following equations:
 - $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$
A: $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$
 - $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
A: $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 8\text{ CO}_2 + 6\text{H}_2\text{O}$
- What isotope is used to define one mole?
A: Carbon-12
- How many particles (atoms or molecules) are in one mole?
A: 6.02×10^{23}
- Find the molar mass for the following chemical formulas:
 - ZnF_2
A: 103.377 g/mol
 - $\text{Al}_2(\text{SO}_4)_3$
A: 342.151 g/mol
- How many moles are in 4.37 kg of NaOH?
A: 109.3 mol
- How many atoms are in 2075 g of He?
A: 3.12×10^{26} atoms
- What volume will 92.31 g of CO₂ occupy at STP?
A: 46.985 L
- How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?
$$\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$$

A: 1.68 g NaOH
- In the reaction, $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$, if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?
A: 126 g MgO

Gases and Their Properties

- What is the kinetic molecular theory (KMT)?
*A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
 - Gas are composed of a smoky particles that behave like hard spherical objects in a state of constant, random motion.
 - These particles move in a straight line until they collide with another particle or the walls of the container.
 - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
 - There is no force of attraction between gas particles or between the particles and the walls of the container.
 - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.*

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J Rink

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Table of Contents Nd Semester Chemistry Exam Review Fot Plato

1. Understanding the eBook Nd Semester Chemistry Exam Review Fot Plato
 - The Rise of Digital Reading Nd Semester Chemistry Exam Review Fot Plato
 - Advantages of eBooks Over Traditional Books
2. Identifying Nd Semester Chemistry Exam Review Fot Plato
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nd Semester Chemistry Exam Review Fot Plato
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nd Semester Chemistry Exam Review Fot Plato

- Personalized Recommendations
- Nd Semester Chemistry Exam Review Fot Plato User Reviews and Ratings
- Nd Semester Chemistry Exam Review Fot Plato and Bestseller Lists

5. Accessing Nd Semester Chemistry Exam Review Fot Plato Free and Paid eBooks

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- Nd Semester Chemistry Exam Review Fot Plato eBook Subscription Services
- Nd Semester Chemistry Exam Review Fot Plato Budget-Friendly Options

6. Navigating Nd Semester Chemistry Exam Review Fot Plato eBook Formats

- ePUB, PDF, MOBI, and More
- Nd Semester Chemistry Exam Review Fot Plato Compatibility with Devices
- Nd Semester Chemistry Exam Review Fot Plato Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Nd Semester Chemistry Exam Review Fot Plato
- Highlighting and Note-Taking Nd Semester Chemistry Exam Review Fot Plato
- Interactive Elements Nd Semester Chemistry Exam Review Fot Plato

8. Staying Engaged with Nd Semester Chemistry Exam Review Fot Plato

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Nd Semester Chemistry Exam Review Fot Plato

9. Balancing eBooks and Physical Books Nd Semester Chemistry Exam Review Fot Plato

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Nd Semester Chemistry Exam Review Fot Plato

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Nd Semester Chemistry Exam Review Fot Plato

- Setting Reading Goals Nd Semester Chemistry Exam Review Fot Plato
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Nd Semester Chemistry Exam Review Fot Plato

- Fact-Checking eBook Content of Nd Semester Chemistry Exam Review Fot Plato
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web from wheat germ protocol salk institute mobile lab step 1 obtain a numbered 50 ml conical vial with 10 ml of water write

your number here step 2 to the tube with 10 ml water add 1g 1 pinch of raw wheat germ the wheat germ should settle to fill 1 2 the cone step 3 cap the tube and vigorously shake to mix for about 30 seconds

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web purpose the purpose of this lab is to extract dna from wheat germ iii hypothesis if you add wheat germ water soap and alcohol then the dna should appear iv materials 1 gram raw wheat germ 50 ml test tube 20 ml hot 50 60 deg c tap water
frac14 teaspoon detergent soap paper towel 14 ml 91 alcohol v procedure 1