

AP Biology Study Guide Unit 1

Properties of Water:

1. Polarity - allows H₂O to form Hydrogen Bonds
2. Cohesion - Water H bonds to itself, creating surface tension
3. Adhesion - Water adheres to other polar substances
 - a. Allows easy transport of water against Gravity
4. Universal Solvent - Water can dissolve Polar and Ionic Compounds
 - a. This allows for reactions between Aqueous Solutions
5. Expands when Frozen - this means that Ice floats
6. High Specific Heat - water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
 - a. High Heat of Vaporization - only warmest molecules can evaporate, so this causes evaporative cooling
7. Can be regulated by non-polar hydrophobic substances
8. Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water

Hydrophilic Substance - Polar substance that attracts water molecules

CHON - Carbon, Hydrogen, Nitrogen, Oxygen

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
 - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
 - Form Stable Covalent Bonds
 - Can form double and triple bonds
 - Versatile

Biological Transition Metals - Mn, Fe, Co, Cu, Zn

- Can exist in multiple oxidation states → can be both oxidized and reduced

Biological Ions: Na⁺, Mg²⁺, Cl⁻, K⁺, Ca²⁺

- Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

5 Foundation Biological Molecules:

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 - The Rise of Digital Reading Apex Bio Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Apex Bio Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Apex Bio Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Apex Bio Study Guide
 - Personalized Recommendations
 - Apex Bio Study Guide User Reviews and Ratings
 - Apex Bio Study Guide and Bestseller Lists
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- 7. Enhancing Your Reading Experience
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 - Interactive Elements Apex Bio Study Guide
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Apex Bio Study Guide
- 9. Balancing eBooks and Physical Books Apex Bio Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Apex Bio Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Apex Bio Study Guide
 - Setting Reading Goals Apex Bio Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Apex Bio Study Guide
 - Fact-Checking eBook Content of Apex Bio Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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it 2 divide the instructions into a series of steps 3 explain each step one by one and 4 use a transition signal to introduce each step