

Balancing Chemical Equations

(What stays the same and what may change in a chemical reaction?)

Model 1: Atoms are conserved in chemical reactions

Atoms are neither created nor destroyed when chemical reactions occur. Two balanced chemical reactions (or equations) are given below. In a chemical reaction equation, reactant atoms are shown on the left side of the arrow and product atoms are shown on the right.



Molecular Picture:



Critical Thinking Questions:

- Considering Equation (1) above:
 - What are the reactants? Write your answer in both words and formulas. Check that your teammates agree.
 - What is the product? Write your team's answer in both words and formulas.
- What does the arrow represent in a chemical reaction?
- How many hydrogen molecules are represented on the reactant side of equation (1)? ____
- How many water molecules are represented on the product side of equation (1)? ____
- How many H atoms are represented on the reactant side of equation (1)? ____
On the product side? ____
- How many O atoms are represented on the reactant side of equation (1)? ____
On the product side? ____
- Compare the number of atoms of each element (H and O) on the reactant and product sides of equation (1). Are these numbers **the same** or **different** (circle one)?
- Compare the number of atoms of each element (Fe, Al, and O) on the reactant and product sides of equation (2). Are these numbers **the same** or **different** (circle one)?
- In reaction (1):
 - For each H_2 molecule consumed, how many H_2O molecules are produced? ____
 - For each O_2 molecule consumed, how many H_2O molecules are produced? ____
 - Recall that a mole is simply a specific large quantity, like a dozen or a gross. For each mole of oxygen molecules consumed, how many moles of water molecules are produced? ____

Answers To Chemactivity 45

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**

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Organic Chemistry, a Guided Inquiry Andrei Straumanis, 2004 Chemistry Richard S. Moog, John J. Farrell, 2017-06-26 In the newly updated 7th Edition Chemistry A Guided Inquiry continues to follow the underlying principles developed by years of extensive research on how students learn and draws on testing by those using the POGIL methodology This text follows the principles of inquiry based learning and correspondingly emphasizes underlying chemistry concepts and the reasoning behind them This text provides an approach that follows modern cognitive learning principles by having students learn how to create knowledge based on experimental data and how to test that knowledge General, Organic, and Biological Chemistry Michael P. Garoutte, 2014-02-24 Classroom activities to support a General Organic and Biological Chemistry text Students can follow a guided inquiry approach as they learn chemistry in the classroom General Organic and Biological Chemistry A Guided Inquiry serves as an accompaniment to a GOB Chemistry text It can suit the one or two semester course This supplemental text supports Process Oriented Guided Inquiry Learning POGIL which is a student focused group learning philosophy of instruction The materials offer ways to promote a student centered science classroom with activities The goal is for students to gain a greater understanding of chemistry through exploration Introductory Chemistry Michael P. Garoutte, Ashley B. Mahoney, 2015-08-10 The ChemActivities found in Introductory Chemistry A Guided Inquiry use the classroom guided inquiry approach and provide an excellent accompaniment to any one semester Introductory text Designed to support Process Oriented Guided Inquiry Learning POGIL these materials provide a variety of ways to promote a student focused active classroom that range from cooperative learning to active student participation in a more traditional setting **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Federation Proceedings*, 1985

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Table of Contents Answers To Chemactivity 45

1. Understanding the eBook Answers To Chemactivity 45
 - The Rise of Digital Reading Answers To Chemactivity 45
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Chemactivity 45
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Chemactivity 45
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Chemactivity 45
 - Personalized Recommendations

- Answers To Chemactivity 45 User Reviews and Ratings
- Answers To Chemactivity 45 and Bestseller Lists
- 5. Accessing Answers To Chemactivity 45 Free and Paid eBooks
 - Answers To Chemactivity 45 Public Domain eBooks
 - Answers To Chemactivity 45 eBook Subscription Services
 - Answers To Chemactivity 45 Budget-Friendly Options
- 6. Navigating Answers To Chemactivity 45 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers To Chemactivity 45 Compatibility with Devices
 - Answers To Chemactivity 45 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Chemactivity 45
 - Highlighting and Note-Taking Answers To Chemactivity 45
 - Interactive Elements Answers To Chemactivity 45
- 8. Staying Engaged with Answers To Chemactivity 45
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Chemactivity 45
- 9. Balancing eBooks and Physical Books Answers To Chemactivity 45
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Chemactivity 45
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers To Chemactivity 45
 - Setting Reading Goals Answers To Chemactivity 45
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers To Chemactivity 45
 - Fact-Checking eBook Content of Answers To Chemactivity 45

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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