

Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: $[\text{Ne}] 3s^2 3p^2$ *may use this form now!!!

2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or $[\text{Kr}] 5s^2 4d^{10} 5p^5$

3) Which family of elements is characterized by an s^2p^3 configuration?

Nitrogen Family, or Group 15

4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a $4s^2 4p^3$ electron configuration **Br, bromine**
- the 4th period Group 16 element **Se, selenium**

5) Identify the elements that have the following outer electron configurations:

- $2s^2 2p^4$ **O, oxygen**
- $4s^2$ **Ca, Calcium**
- $4s^2 3d^{10}$ **Zn, zinc**

6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

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Gardner

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Chemistry David W. Oxtoby,Wade A. Freeman,Toby F. Block,1998 Textbook of chemistry concepts

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Student Study Guide to accompany Chemistry Martin Silberberg,Libby Weberg,2005-01-06

Engineer-in-training License Review Donald G. Newnan,1995

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Table of Contents 14 Chemical Periodicity Practice Problems Answers

1. Understanding the eBook 14 Chemical Periodicity Practice Problems Answers
 - The Rise of Digital Reading 14 Chemical Periodicity Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 14 Chemical Periodicity Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 14 Chemical Periodicity Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 14 Chemical Periodicity Practice Problems Answers
 - Personalized Recommendations

- 14 Chemical Periodicity Practice Problems Answers User Reviews and Ratings
- 14 Chemical Periodicity Practice Problems Answers and Bestseller Lists

5. Accessing 14 Chemical Periodicity Practice Problems Answers Free and Paid eBooks

- 14 Chemical Periodicity Practice Problems Answers Public Domain eBooks
- 14 Chemical Periodicity Practice Problems Answers eBook Subscription Services
- 14 Chemical Periodicity Practice Problems Answers Budget-Friendly Options

6. Navigating 14 Chemical Periodicity Practice Problems Answers eBook Formats

- ePUB, PDF, MOBI, and More
- 14 Chemical Periodicity Practice Problems Answers Compatibility with Devices
- 14 Chemical Periodicity Practice Problems Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 14 Chemical Periodicity Practice Problems Answers
- Highlighting and Note-Taking 14 Chemical Periodicity Practice Problems Answers
- Interactive Elements 14 Chemical Periodicity Practice Problems Answers

8. Staying Engaged with 14 Chemical Periodicity Practice Problems Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 14 Chemical Periodicity Practice Problems Answers

9. Balancing eBooks and Physical Books 14 Chemical Periodicity Practice Problems Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 14 Chemical Periodicity Practice Problems Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 14 Chemical Periodicity Practice Problems Answers

- Setting Reading Goals 14 Chemical Periodicity Practice Problems Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 14 Chemical Periodicity Practice Problems Answers

- Fact-Checking eBook Content of 14 Chemical Periodicity Practice Problems Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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