

ATOMIC STRUCTURE

Particle	Symbol	Charge	Relative charge	Mass	Relative mass w.r.t H	$\frac{e}{m}$ ratio
Electron	e^-	$-1.6 \times 10^{-19} \text{ C}$	-1	$9.1 \times 10^{-28} \text{ gram}$	$1/1837$	1837
Proton	p^+	$+1.6 \times 10^{-19} \text{ C}$	+1	$1.672 \times 10^{-24} \text{ g}$	1	1
Neutron	n^0	Zero	Zero	$1.672 \times 10^{-24} \text{ g}$	1	0
α -Particle	${}_2\text{He}^4$	$2 \times 1.6 \times 10^{-19} \text{ C}$	+2	$4 \times 1.672 \times 10^{-24} \text{ g}$	4	$1/2$

Thomson's atomic model

(Plum pudding / Watermelon model)

- +ve charge $\rightarrow$ uniformly distributed
- -ve charge $\rightarrow$ embedded in cloud of the charge.
- Limitation $\rightarrow$ doesn't explain stability

Rutherford's experiment

(Discovery of nucleus)

Atom $\rightarrow$ Nucleus (Proton)
 $\rightarrow$ Extra nuclear part (e^-)

$$R_{\text{nucleus}} = K(A)^{1/3}$$

$$\left(\frac{R_{\text{nuc}}}{A}\right)_1 = \left(\frac{A_1}{A_2}\right)^{1/3}$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi R^3$$

$$\frac{V_{\text{nuc1}}}{V_{\text{nuc2}}} = \frac{A_1}{A_2}$$

$$V_{\text{atom}} \approx 10^{15} \times V_{\text{nucleus}}$$

Term associated with element

Z $\swarrow$ $\searrow$
 $A \rightarrow$ Mass number ($P+n$)
 $\rightarrow$ Atomic number (P)

- **Isotope** $\rightarrow Z \rightarrow$ Same $A \rightarrow$ different
- **Isobar** $\rightarrow Z \rightarrow$ different $A \rightarrow$ same
- **Isotone** $\rightarrow A - Z = n \rightarrow$ same
- **Isotectonic** $\rightarrow$ Same no. of e^- (NO^0, CO)
- **Isosteric** $\rightarrow$ Same no. of atom and e^- (CO_2, N_2)
- **Isotopomers** $\rightarrow (n-p) \text{ or } (A-2Z) \rightarrow$ same
- **Paramagnetic** $\rightarrow$ Having non zero unpaired e^-
- **Diamagnetic** $\rightarrow$ Having zero unpaired e^-

Dual nature of light

Wave nature

• Wave length (λ)
 distance between crest or trough or length of one wave.

• Wave no. $\boxed{\bar{\nu} = \frac{1}{\lambda}}$

• frequency $\boxed{\nu = \frac{c}{\lambda}}$

$c \rightarrow 3 \times 10^8 \text{ m/s}$

$\nu \propto \frac{1}{\lambda} \rightarrow \boxed{\frac{\nu_1}{\nu_2} = \frac{\lambda_2}{\lambda_1}}$

Particle nature

• photon $\rightarrow$ quanta of light.

$\boxed{E = h\nu}$

$h \rightarrow$ plank's constant
 $h = 6.6 \times 10^{-34} \text{ joule sec}$

$\boxed{E = n \times E_{\text{min}}}$

$\downarrow$
 integral multiple

$\boxed{E = h\nu = \frac{hc}{\lambda}}$

$\boxed{\frac{E_1}{E_2} = \frac{\nu_1}{\nu_2}} \quad \boxed{\frac{E_1}{E_2} = \frac{\lambda_2}{\lambda_1}}$

Photo electric effect

- example of particle nature.
- Surface phenomena
- Single photon release single e^-

$\boxed{h\nu = \omega + KE} \quad \omega \rightarrow \text{work function}$

$\boxed{h\nu = h\nu_0 + \frac{1}{2} m_e v_e^2} \quad \nu \rightarrow \text{incident frequency of photon}$

$\boxed{\frac{hc}{\lambda} = \frac{hc}{\lambda_0} + \frac{1}{2} m_e v_e^2} \quad \nu_0 \rightarrow \text{Threshold frequency of metal.}$

$\nu > \nu_0 \rightarrow e^-$ emitted with KE
 $(\lambda < \lambda_0)$

$\nu = \nu_0 \rightarrow e^-$ emitted without KE
 $(\lambda = \lambda_0)$

$\nu < \nu_0 \rightarrow e^-$ not emitted.
 $(\lambda > \lambda_0)$

Atomic Structure Quick Xword Answers

Guy Murchie



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