

Use your notes from the Atomic Structure program to answer the following questions.

1. The atomic number tells the number of positively charged protons in the nucleus of an atom. The atom is neutral because this is also the number of negatively charged electrons in the atom.
2. The mass number tells the total number of protons and neutrons in the nucleus of an atom. These particles collectively are called nucleons since both are located in the nucleus.
3. Isotopes are atoms of the same element with different numbers of neutrons which results in different mass numbers.
4. Write the hyphen notation of the three isotopes of hydrogen:
  - a. Hydrogen - 1  $^1\text{H}$
  - b. Hydrogen - 2  $^2\text{H}$
  - c. Hydrogen - 3  $^3\text{H}$
5. Fill in the following chart (Hint: Isotopes may have different masses than the periodic table lists!):

Symbol	Atomic #	Mass #	# of protons	# of electrons	# of neutrons
Hg	80	201	80	80	121
Mo	42	96	42	42	54
Cu	29	64	29	29	35
C	6	14	6	6	8
	15	31	15	15	16
Pb	82	208	82	82	126
Na	11	24	11	11	13

## Atomic Structure 4 Answers

**Yan Bai**

## Atomic Structure 4 Answers:

*Atomic Structure and Chemical Bond: A Problem Solving Approach* Manas Chandra, 2019-05-03 particle in a box and to the hydrogen atom quantization of energy levels uncertainty principle probability distribution functions angular and radial wave functions nodal properties sectional and charge cloud representation of atomic orbitals etc have been covered in detail The valence bond and molecular orbital methods of bonding hybridization orbital structure of common hydrocarbons bonding in coordination compounds based on valence bond and ligand field theories the concept of valency ionic and covalent bonding bonding in metals secondary bond forces and so on have been discussed in a reasonable amount of detail A unique feature of the book is the adoption of a problem solving approach Thus while the text has been frequently interspersed with numerous fully worked out illustrative examples to help the concepts and theories a large number of fully solved problems have been appended at the end of each chapter totalling nearly 300 With its lucid style and in depth coverage the book would be immensely useful to undergraduate and postgraduate students of general chemistry and quantum chemistry Students of physics and materials science would also find the book an invaluable supplement

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