

Name Key Class _____ Date _____

Atomic Structure

Subatomic Particles in Neutral Atoms

The table below contains information about several isotopes. Use the information given to fill in the blanks. Assume all atoms are neutral.

Isotope Name	Nuclear Symbol	Atomic Number	Mass Number	# of Protons	# of Neutrons	# of Electrons
1. calcium-40	$^{40}_{20} Ca$	20	40	20	20	20
2. iron-56	$^{56}_{26} Fe$	26	56	26	30	26
3. oxygen-18	$^{18}_{8} O$	8	18	8	10	8
4. gold-197	$^{197}_{79} Au$	79	197	79	118	79

Subatomic Particles in Ions

The table below contains information about several ions. Use the information given to fill in the blanks.

Element Name	Ion Symbol	Atomic Number	Mass Number	# of Protons	# of Neutrons	# of Electrons
5. chlorine	Cl^{-}	17	35	17	18	18
6. silver	Ag^{+}	47	107	47	60	46
7. oxygen	O^{2-}	8	16	8	8	10
8. aluminum	Al^{3+}	13	27	13	14	10

Average Atomic Mass

9. Calculate the average atomic mass for neon if its abundance in nature is 90.5% neon-20, 0.3% neon-21, and 9.2% neon-22.

$$\frac{(90.5)(20 \text{ amu}) + (0.3)(21 \text{ amu}) + (9.2)(22 \text{ amu})}{100} = 20.18 \text{ amu}$$

10. Calculate the average atomic mass of silver if 13 out of 25 atoms are silver-107 and 12 out of 25 atoms are silver-109.

$$\frac{(13)(107 \text{ amu}) + (12)(109 \text{ amu})}{25} = 107.96 \text{ amu}$$

Atomic Mass And Answer Key

Larry D. Kirkpatrick, Gerald F. Wheeler

Atomic Mass And Answer Key:

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Table of Contents Atomic Mass And Answer Key

1. Understanding the eBook **Atomic Mass And Answer Key**
 - The Rise of Digital Reading **Atomic Mass And Answer Key**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Atomic Mass And Answer Key**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Atomic Mass And Answer Key**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Atomic Mass And Answer Key**
 - Personalized Recommendations
 - **Atomic Mass And Answer Key** User Reviews and Ratings
 - **Atomic Mass And Answer Key** and Bestseller Lists
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- Atomic Mass And Answer Key Budget-Friendly Options

6. Navigating Atomic Mass And Answer Key eBook Formats

- ePUB, PDF, MOBI, and More
- Atomic Mass And Answer Key Compatibility with Devices
- Atomic Mass And Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Atomic Mass And Answer Key
- Highlighting and Note-Taking Atomic Mass And Answer Key
- Interactive Elements Atomic Mass And Answer Key

8. Staying Engaged with Atomic Mass And Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Atomic Mass And Answer Key

9. Balancing eBooks and Physical Books Atomic Mass And Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Atomic Mass And Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Atomic Mass And Answer Key

- Setting Reading Goals Atomic Mass And Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Atomic Mass And Answer Key

- Fact-Checking eBook Content of Atomic Mass And Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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