

Student Exploration: Free-Cell Laboratory

Objective: Understand the relationship between Gibbs free energy, enthalpy, and entropy.

Background: Gibbs free energy is a thermodynamic potential that is used to predict the spontaneity of a process.

1. Gibbs free energy is defined as the total energy of a system minus the product of temperature and entropy. It is denoted by G .

2. The change in Gibbs free energy for a process is denoted by ΔG . It is calculated as $\Delta G = \Delta H - T\Delta S$.

3. If ΔG is negative, the process is spontaneous. If ΔG is positive, the process is non-spontaneous. If ΔG is zero, the system is at equilibrium.

Example 1:

Consider the reaction: $2H_2(g) + O_2(g) \rightarrow 2H_2O(l)$. The standard enthalpy of formation for $H_2O(l)$ is -285.8 kJ/mol . The standard entropy of formation for $H_2O(l)$ is $69.9 \text{ J/mol}\cdot\text{K}$.

4. Calculate ΔG for the reaction at 25°C .

5. Is the reaction spontaneous?

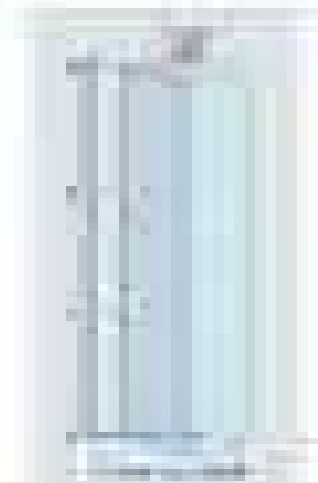
6. Calculate ΔG for the reaction at 100°C .

7. Is the reaction spontaneous at 100°C ?

8. Calculate ΔG for the reaction at 125°C .

9. Calculate ΔG for the reaction at 150°C .

10. Calculate ΔG for the reaction at 175°C .



Answer Key To The Fall Laboratory Gizmo

O García



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Answer Key To The Fall Laboratory Gizmo Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Answer Key To The Fall Laboratory Gizmo**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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