

Chemistry

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(Chapter 1) (Solutions) (Intext Questions)
(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C_6H_6) and carbon tetrachloride (CCl_4) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

$$\text{Mass percentage of } C_6H_6 = \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solution}} \times 100\% = \frac{\text{Mass of } C_6H_6}{22 + 122} \times 100\%$$

$$= \frac{22}{22 + 122} \times 100\% = 15.28\%$$

$$\text{Mass percentage of } CCl_4 = \frac{\text{Mass of } CCl_4}{\text{Total mass of the solution}} \times 100\% = \frac{\text{Mass of } CCl_4}{22 + 122} \times 100\%$$

$$= \frac{122}{22 + 122} \times 100\% = 84.72\%$$

Alternatively, Mass percentage of CCl_4 = $(100 - 15.28)\% = 84.72\%$

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

∴ Mass of carbon tetrachloride = $(100 - 30)$ g = 70 g

Molar mass of benzene (C_6H_6) = $(6 \times 12 + 6 \times 1)$ g mol^{-1} = 78 g mol^{-1}

∴ Number of moles of $C_6H_6 = \frac{30}{78}$ mol = 0.3846 mol

Molar mass of carbon tetrachloride (CCl_4) = $1 \times 12 + 4 \times 35.5$
= 154 g mol^{-1}

∴ Number of moles of $CCl_4 = \frac{70}{154}$ mol = 0.4545 mol

Thus, the mole fraction of C_6H_6 is given as:

$$\begin{aligned} &= \frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4} \\ &= \frac{0.3846}{0.3846 + 0.4545} \\ &= 0.458 \end{aligned}$$

Question 1.3:

Calculate the molarity of each of the following solutions:

(a) 30 g of $Co(NO_3)_2 \cdot 6H_2O$ in 4.3 L of solution

(b) 30 mL of 0.5 M H_2SO_4 diluted to 500 mL

Answer 1.3:

Molarity is given by: $Molarity = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$

(a) Molar mass of $Co(NO_3)_2 \cdot 6H_2O = 59 + 2(14 + 3 \times 16) + 6 \times 18 = 291$ g mol^{-1}

∴ Moles of $Co(NO_3)_2 \cdot 6H_2O = \frac{30}{291}$ mol = 0.103 mol

Therefore, molarity = $\frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$

(b) Number of moles present in 1000 mL of 0.5 M $H_2SO_4 = 0.5 \text{ mol}$

∴ Number of moles present in 30 mL of 0.5 M $H_2SO_4 = \frac{0.5 \times 30}{1000}$ mol = 0.015 mol

Therefore, molarity = $\frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$

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Solutions Manual to Accompany Elements of Physical Chemistry David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book. The manual provides helpful comments and friendly advice to aid understanding. It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions. *Longman Science Chemistry 9* Kohli Nitin, 2008-09

Chapter-wise NCERT + Exemplar + Practice

Questions with Solutions for CBSE Chemistry Class 11 - 2nd Edition Disha Experts, 2017-08-29 The book Chapter wise NCERT Exemplar Practice Questions with Solutions for CBSE Class 11 Chemistry has been divided into 3 parts. Part A provides detailed solutions Question by Question of all the questions/exercises provided in the NCERT Textbook. Part B provides solutions to the questions in the NCERT Exemplar book. Part C provides selected Practice Questions useful for the Class 11 examination along with detailed solutions. The solutions have been designed in such a manner Step by Step that it would bring 100% Concept Clarity for the student.

IIT JEE Chemistry Exam Leaders Expert, Prentice Hall Physical

Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day. Relevant content, lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them. Now includes even more technology tools and activities to support differentiated instruction.

Fundamentals of Environmental Chemistry, Third Edition Stanley E. Manahan, 2011-03-05

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Resources in Education
,1977 **Fundamentals of Sustainable Chemical Science** Stanley E. Manahan,2009-03-10 Written by Stanley Manahan

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Table of Contents 8 Chemistry Sol Answers

1. Understanding the eBook 8 Chemistry Sol Answers
 - The Rise of Digital Reading 8 Chemistry Sol Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 8 Chemistry Sol Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 8 Chemistry Sol Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 8 Chemistry Sol Answers
 - Personalized Recommendations
 - 8 Chemistry Sol Answers User Reviews and Ratings
 - 8 Chemistry Sol Answers and Bestseller Lists
5. Accessing 8 Chemistry Sol Answers Free and Paid eBooks
 - 8 Chemistry Sol Answers Public Domain eBooks
 - 8 Chemistry Sol Answers eBook Subscription Services

- 8 Chemistry Sol Answers Budget-Friendly Options
- 6. Navigating 8 Chemistry Sol Answers eBook Formats
 - ePUB, PDF, MOBI, and More
 - 8 Chemistry Sol Answers Compatibility with Devices
 - 8 Chemistry Sol Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 8 Chemistry Sol Answers
 - Highlighting and Note-Taking 8 Chemistry Sol Answers
 - Interactive Elements 8 Chemistry Sol Answers
- 8. Staying Engaged with 8 Chemistry Sol Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 8 Chemistry Sol Answers
- 9. Balancing eBooks and Physical Books 8 Chemistry Sol Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 8 Chemistry Sol Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 8 Chemistry Sol Answers
 - Setting Reading Goals 8 Chemistry Sol Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 8 Chemistry Sol Answers
 - Fact-Checking eBook Content of 8 Chemistry Sol Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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