

# Chemistry

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(Chapter 1)(Solutions) (Intext Questions)

(Class 12)

## Question 1.1:

Calculate the mass percentage of benzene ( $C_6H_6$ ) and carbon tetrachloride ( $CCl_4$ ) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

### Answer 1.1:

$$\text{Mass percentage of } C_6H_6 = \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solutions}} \times 100 \%$$

$$= \frac{\text{Mass of } C_6H_6}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{22}{22 + 122} \times 100 \% = 15.28 \%$$

$$\text{Mass percentage of } CCl_4 = \frac{\text{Mass of } CCl_4}{\text{Total mass of the solutions}} \times 100 \%$$

$$= \frac{\text{Mass of } CCl_4}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{122}{22 + 122} \times 100 \% = 84.72 \%$$

$$\text{Alternatively, Mass percentage of } CCl_4 = (100 - 15.28) \% = 84.72 \%$$

## Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

### Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

$$\therefore \text{Mass of carbon tetrachloride} = (100 - 30) \text{ g} = 70 \text{ g}$$

$$\text{Molar mass of benzene } (C_6H_6) = (6 \times 12 + 6 \times 1) \text{ g mol}^{-1} = 78 \text{ g mol}^{-1}$$

$$\therefore \text{Number of moles of } C_6H_6 = \frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$$

$$\text{Molar mass of carbon tetrachloride } (CCl_4) = 1 \times 12 + 4 \times 35.5 \\ = 154 \text{ g mol}^{-1}$$

$$\therefore \text{Number of moles of } CCl_4 = \frac{70}{154} \text{ mol} = 0.4545 \text{ mol}$$

Thus, the mole fraction of  $C_6H_6$  is given as:

$$= \frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4}$$

$$= \frac{0.3846}{0.3846 + 0.4545}$$

$$= 0.458$$

## Question 1.3:

Calculate the molarity of each of the following solutions:

(a) 30 g of  $Co(NO_3)_2 \cdot 6H_2O$  in 4.3 L of solution

(b) 30 mL of 0.5 M  $H_2SO_4$  diluted to 500 mL

### Answer 1.3:

$$\text{Molarity is given by: } \text{Molarity} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$$

$$\text{(a) Molar mass of } Co(NO_3)_2 \cdot 6H_2O = 59 + 2(14 + 3 \times 16) + 6 \times 18 = 291 \text{ g mol}^{-1}$$

$$\therefore \text{Moles of } Co(NO_3)_2 \cdot 6H_2O = \frac{30}{291} \text{ mol} = 0.103 \text{ mol}$$

$$\text{Therefore, molarity} = \frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$$

$$\text{(b) Number of moles present in 1000 mL of 0.5 M } H_2SO_4 = 0.5 \text{ mol}$$

$$\therefore \text{Number of moles present in 30 mL of 0.5 M } H_2SO_4 = \frac{0.5 \times 30}{1000} \text{ mol} = 0.015 \text{ mol}$$

$$\text{Therefore, molarity} = \frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$$

## 8 Chemistry Sol Answers

**O García**



## 8 Chemistry Sol Answers:

Solutions Manual to Accompany Elements of Physical Chemistry David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions *Longman Science Chemistry 9* Kohli Nitin, 2008-09 **Chapter-wise NCERT + Exemplar + Practice Questions with Solutions for CBSE Chemistry Class 11 - 2nd Edition** Disha Experts, 2017-08-29 The book Chapter wise NCERT Exemplar Practice Questions with Solutions for CBSE Class 11 Chemistry has been divided into 3 parts Part A provides detailed solutions Question by Question of all the questions exercises provided in the NCERT Textbook Part B provides solutions to the questions in the NCERT Exemplar book Part C provides selected Practice Questions useful for the Class 11 examination along with detailed solutions The solutions have been designed in such a manner Step by Step that it would bring 100% Concept Clarity for the student *IIT JEE Chemistry Exam Leaders Expert*, **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Fundamentals of Environmental Chemistry, Third Edition** Stanley E. Manahan, 2011-03-05 Written by an expert using the same approach that made the previous two editions so successful Fundamentals of Environmental Chemistry Third Edition expands the scope of book to include the strongly emerging areas broadly described as sustainability science and technology including green chemistry and industrial ecology The new edition includes Increased emphasis on the applied aspects of environmental chemistry Hot topics such as global warming and biomass energy Integration of green chemistry and sustainability concepts throughout the text More and updated questions and answers including some that require Internet research Lecturers Pack on CD ROM with solutions manual PowerPoint presentations and chapter figures available upon qualifying course adoptions The book provides a basic course in chemical science including the fundamentals of organic chemistry and biochemistry The author uses real life examples from environmental chemistry green chemistry and related areas while maintaining brevity and simplicity in his explanation of concepts Building on this foundation the book covers environmental chemistry broadly defined to include sustainability aspects green chemistry industrial ecology and related areas These chapters are organized around the five environmental spheres the hydrosphere atmosphere geosphere biosphere and the anthrosphere The last two chapters discuss analytical

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,1977 **Fundamentals of Sustainable Chemical Science** Stanley E. Manahan,2009-03-10 Written by Stanley Manahan

Fundamentals of Sustainable Chemical Science has been carefully designed to provide a basic introduction to chemistry including organic chemistry and biochemistry for readers with little or no prior background in the subject Manahan

bestselling author of many environmental texts presents the material in a practical Solutions to Learning Elementary

Chemistry for Class 8 Dr. R. Goel,Goyal Brothers Prakashan,2021-01-01 Goyal Brothers Prakashan The Chemical News :

and Journal of Physical Science ,1871 *Chemical News and Journal of Physical Science* ,1871

**Fundamentals of Physical Chemistry** Ananya Ganguly,2011 Fundamentals of Physical Chemistry is the signature compilation of the class

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**Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed** Paul A. Bartlett,Judith G. Koch,1992

A Reader's Guide to Contemporary Literature William Swan Sonnenschein,1895 Chemical News and Journal of Industrial Science ,1893

The Pharmaceutical Era ,1894 **The Chemical News** ,1871 *The Chemical*

*News and Journal of Industrial Science* William Crookes,James H. Gardiner,Gerald Druce,H. W. Blood-Ryan,1877 **The**

**Pharmaceutical Era** Charles W. Parsons,1894 Intermediate Algebra Tom Carson,Ellyn Gillespie,Bill E. Jordan,2005

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due to bad experiences in past math courses Developmental students often have never developed nor applied a study system in mathematics To address these needs the authors have framed three goals for Intermediate Algebra 1 reduce math anxiety

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