

1989 Advanced Placement Exam

Part I Multiple Choice

Note: For all questions referring to solutions, assume that the solvent is water unless otherwise stated.

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then blacken the corresponding space on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–3

- a. O c. Rb e. N
b. La d. Mg

- What is the most electronegative element?
- Which element exhibits the greatest number of different oxidation states?
- Which of the elements above has the smallest ionic radius for its most commonly found ion?

Questions 4–7

- a. $1s^2 2s^2 2p^5 3s^2 3p^5$
b. $1s^2 2s^2 2p^6 3s^2 3p^6$
c. $1s^2 2s^2 2p^6 2d^{10} 3s^2 3p^6$
d. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5$
e. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^3 4s^2$

- An impossible electronic configuration
- The ground-state configuration for the atoms of a transition element
- The ground-state configuration of a negative ion of a halogen
- The ground-state configuration of a common ion of an alkaline earth element

Questions 8–10

- a. a solution with a pH less than 7 that is not a buffer solution
b. a buffer solution with a pH between 4 and 7
c. a buffer solution with a pH between 7 and 10
d. a solution with a pH greater than 7 that is not a buffer solution
e. a solution with a pH of 7

(Ionization constants:



- A solution prepared to be initially 1 M in NaCl and 1 M in HCl.
- A solution prepared to be initially 1 M in Na_2CO_3 and 1 M in CH_3COONa
- A solution prepared to be initially 0.5 M in CH_3COOH and 1 M in CH_3COONa

Questions 11–14

- a. hydrogen bonding
b. hybridization
c. ionic bonding
d. resonance
e. van der Waals forces (London dispersion forces)

- Is used to explain why iodine molecules are held together in the solid state
- Is used to explain why the boiling point of HF is greater than the boiling point of HBr
- Is used to explain the fact that the four bonds in methane are equivalent
- Is used to explain the fact that the carbon-to-carbon bonds in benzene, C_6H_6 , are identical

Directions: Each of the questions or incomplete statements below is followed by five suggested answers or compilations. Select the one that is best in each case and then blacken the corresponding space on the answer sheet.

2008 Chemistry Ap Multiple Choice

Rachel S Tattersall



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Developing and Validating Test Items Thomas M. Haladyna, Michael C. Rodriguez, 2013-07-18 Since test items are the building blocks of any test learning how to develop and validate test items has always been critical to the teaching learning process As they grow in importance and use testing programs increasingly supplement the use of selected response multiple choice items with constructed response formats This trend is expected to continue As a result a new item writing book is needed one that provides comprehensive coverage of both types of items and of the validity theory underlying them This book is an outgrowth of the author's previous book *Developing and Validating Multiple Choice Test Items* 3e Haladyna 2004 That book achieved distinction as the leading source of guidance on creating and validating selected response test items Like its predecessor the content of this new book is based on both an extensive review of the literature and on its author's long experience in the testing field It is very timely in this era of burgeoning testing programs especially when these items are delivered in a computer based environment Key features include Comprehensive and Flexible No other book so thoroughly covers the field of test item development and its various applications Focus on Validity Validity the most important consideration in testing is stressed throughout and is based on the Standards for Educational and Psychological Testing currently under revision by AERA APA and NCME Illustrative Examples The book presents various selected and constructed response formats and uses many examples to illustrate correct and incorrect ways of writing items Strategies for training item writers and developing large numbers of items using algorithms and other item generating methods are also presented Based on Theory and Research A comprehensive review and synthesis of existing research runs throughout the book and complements the expertise of its authors SAGE Handbook of Research on Classroom Assessment James H.

McMillan, 2013 The Sage Handbook of Research on Classroom Assessment provides scholars professors graduate students and other researchers and policy makers in the organizations agencies testing companies and school districts with a comprehensive source of research on all aspects of K 12 classroom assessment The handbook emphasizes theory conceptual frameworks and all varieties of research quantitative qualitative mixed methods to provide an in depth understanding of the knowledge base in each area of classroom assessment and how to conduct inquiry in the area It presents classroom assessment research to convey in depth the state of knowledge and understanding that is represented by the research with particular emphasis on how classroom assessment practices affect student achievement and teacher behavior Editor James H. McMillan and five Associate Editors bring the best thinking and analysis from leading classroom assessment researchers on the nature of the research making significant contributions to this prominent and hotly debated topic in education **AP** Philip Michael Sadler, 2010 Draws together the most recent and rigorous research on the strengths and weaknesses of the Advanced Placement program Examines closely the differences between AP and other high school courses as well as variations among AP courses In depth studies gauge the impact of AP coursework on student performance in college Finally

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