

Activation Wizard

Follow these steps to activate your software over the telephone.

Step 1: Select the country/region you are calling from and call the Product Activation Center using any of the telephone numbers provided.

select a country/region

▼

Step 2: When prompted, provide this Installation Code:
012345-012345-012345-012345-012345-012345-012345-012345

Step 3: Enter your Confirmation Code here:

A	<input type="text"/>	B	<input type="text"/>	C	<input type="text"/>	D	<input type="text"/>	E	<input type="text"/>	F	<input type="text"/>	G	<input type="text"/>
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Disha Experts



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web in this article we ll look at how we can use the stoichiometric relationships contained in balanced chemical equations to determine amounts of substances consumed and produced in chemical reactions balanced equations and mole ratios

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web jul 4 2019 solution from the balanced equation it is known that 1 mol N_2 2 mol NH_3 use the periodic table to look at the atomic weights of the elements and calculate the weights of the reactants and products 1 mol of N_2 28.0 g 1 mol of NH_3 is 17.0 g

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web solution average atomic mass of Li 6.941 amu $\times 6.0151$ amu $\times 7.0160$ amu 6.941 1.0009 $\times 7.0160$ 1.0009 $\times 0.075$ $\times 0.075$ corresponds to a natural abundance of 6Li of 7.5 percent the natural abundance of 7Li is $\times 0.925$ or 92.5 percent 3.7 the unit factor required is 6.022

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web a molar mass Li_2CO_3 73.89 g CS_2 76.15 g c molar mass CHCl_3 119.37 g d molar mass $\text{C}_6\text{H}_8\text{O}_6$ 176.12 g e molar mass KNO_3 101.11 g

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web 41.42 chemistry ch 3 mass relationships in chemical reactions molecular mass of NO_2 atomic mass of N 14 atomic mass of

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108 02 amu example 3 2 exercise 3 3