

112. The IUPAC name of



is

- (1) 3, 1-diethyl-4,2,2-dimethylpentane  
(2) 4, 4-dimethyl-5, 5-diethylpentane  
(3) 5, 5-diethyl-4, 4-dimethylpentane  
(4) 3-ethyl-4, 4-dimethylheptane

Ans.  
Sol.



The correct answer is 3-ethyl-4, 4-dimethylheptane.

Hence, (4) is correct.

113. Which of the following species exhibits the diamagnetic behaviour?

- (1)  $\text{O}_2^+$   
(2)  $\text{O}_2$   
(3)  $\text{O}_2^-$   
(4)  $\text{NO}$

Ans.  
Sol.

The correct option is  $\text{O}_2^+$

This species has 18  $\sigma$ , which are filled in such a way that all molecular orbitals are fully filled, so diamagnetic.



Hence, (1) is correct.

114. The stability of divalides of Sn, Ge, Si and Pb increases steadily in the sequence

- (1)  $\text{GeX}_2 < \text{SnX}_2 < \text{SnX}_2 < \text{PbX}_2$   
(2)  $\text{SnX}_2 < \text{GeX}_2 < \text{PbX}_2 < \text{SnX}_2$   
(3)  $\text{SnX}_2 < \text{GeX}_2 < \text{SnX}_2 < \text{PbX}_2$   
(4)  $\text{PbX}_2 < \text{SnX}_2 < \text{GeX}_2 < \text{SnX}_2$

Ans.  
Sol.

Due to inert pair effect, the stability of +2 oxidation state increases as we move down this group.  
 $\therefore \text{SnX}_2 < \text{GeX}_2 < \text{SnX}_2 < \text{PbX}_2$

Hence, (3) is correct.

115. Identify the incorrect statement among the following

- (1) Ozone reacts with  $\text{SO}_2$  to give  $\text{SO}_3$   
(2) Silicon reacts with  $\text{NaOH}_{\text{conc}}$  in the presence of air to give  $\text{Na}_2\text{SiO}_3$  and  $\text{H}_2\text{O}$   
(3)  $\text{Cl}_2$  reacts with excess of  $\text{NH}_3$  to give  $\text{N}_2$  and  $\text{HCl}$   
(4)  $\text{Br}_2$  reacts with hot and strong  $\text{NaOH}$  solution to give  $\text{NaBr}$ ,  $\text{NaBrO}_3$  and  $\text{H}_2\text{O}$

Ans.  
Sol.

(4)  $\text{Br}_2$  reacts with hot and strong  $\text{NaOH}$  to give  $\text{NaBr}$ ,  $\text{NaBrO}_2$  and  $\text{H}_2\text{O}$ .

Hence, (4) is incorrect statement.

116. The charge/radius ratio of a cation determines its polarizing power. Which one of the following sequences represents the increasing order of the polarizing power of the cationic species,  $\text{K}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Be}^{2+}$ ?

- (1)  $\text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}$   
(2)  $\text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}$   
(3)  $\text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}$   
(4)  $\text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+$

Ans.  
Sol.

(3) Higher the charge/radius ratio, more is the polarizing power.



Hence, (3) is correct.

117. The density ( $\text{in g mL}^{-1}$ ) of a 3.65 M sulphuric acid solution that is 20%  $\text{H}_2\text{SO}_4$  (Molar mass = 98 g  $\text{mol}^{-1}$ ) by mass will be

- (1) 1.64  
(2) 1.88  
(3) 1.22  
(4) 1.45

Ans.  
Sol.

Let the density of solution be  $d$

Molarity of solution given = 3.6

i.e. 1 litre of solution contains 3.6 moles of  $\text{H}_2\text{SO}_4$

or 1 litre of solution contains  $3.6 \times 98$  gms of  $\text{H}_2\text{SO}_4$

Since, the solution is 20% by mass,

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