

AP Chemistry

Gas Laws and Stoichiometry

3 Scenarios:

1. If the P and T in an all gas system remain constant and the problem is one of volume going to volume...
 - **Gay-Lussac's Law** – If the T and P in an all gas system remain constant, the volumes of reacting gases are directly related to the coefficients of the balanced system.



- Eliminates having to find moles in the problem, thus...
- Eliminates having to use molar volume or $PV = nRT$ in the problem

Example: What volume of hydrogen gas would combine with 4.2 L. of oxygen gas to form steam at a constant 750 °C and atmospheric pressure?

2. If conditions are at STP but there are grams involved...

- **Molar Volume** – Convert the grams to moles, use the balanced equation, and then use molar volume to find the volume of gas produced.

Example: Calculate the liters of CO_2 measured at 0°C and 1 atm that could be obtained by allowing 45.0 g. of calcium carbonate to react with an excess of hydrochloric acid (The molar mass of calcium carbonate is 100 g/mol).

Ap Chemistry Stoichiometry Study Guide

Mary Millhollon, Richard H. Langley

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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end. Two full length practice tests with detailed answer explanations are included in the book.

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