

Chapter 2 Review

1. Which properties are characteristics of metals?
Which property means that you can punch the substance into a ball?

(a) ductility (b) malleability
(c) conductivity (d) brittleness

2. Which of the following is a transition metal?
(a) Cu (b) H (c) Pb (d) He

3. Which of the following is a transition metal?
(a) Cu (b) P (c) Cu (d) U

4. Which of the following are alkali metals?
(a) Mg (b) Na (c) Si (d) Al (e) Li

5. Which of the following is an halogen?
(a) Se (b) Br (c) Cl (d) F (e) W

6. Which element has the highest melting point?
(a) Pb (b) Au (c) Cu (d) Mg
(e) Ca

7. Which part is part of the
negative electrode? (a) positive electrode
(b) power source (c) gas inside the tube

8. In a cathode ray tube electrons are held toward
(a) a positively charged plate
(b) a negatively charged plate

9. Listed below are the charges and masses of four
isotopes. Which one will be deflected by 90° in a
mass spectrometer?
(a) 17.0 mass
(b) 18.0 mass (c) 19.0 mass

Name _____
Period _____ Date _____

10. In a addition of long type expression, the charge
on that of an isotope of Carbon has to be 10.5.

Isotope	Carbon-12	1.00	Carbon-13	1.01
Isotope	Carbon-14	1.00	Carbon-15	1.01
Isotope	Carbon-16	1.00	Carbon-17	1.01

What is the charge of the electron according to this
expression?
(a) 1.01 Coulomb (b) 1.00 Coulomb
(c) 1.02 Coulomb (d) 1.000 Coulomb

11. Picture below is a schematic of the Rutherford
experiment. Which numbered particle gives the
best evidence for the nuclear theory?



12. Which of the following is an isotope of the element
with 20 protons (20 protons = 20 neutrons = 20)?
(a) Element 20 (b) Element 21
(c) isotope 20 (d) isotope 21

13. The radioactive element X has the following natural
abundance and average masses. What is the mass
of X?
(a) 17.00 mass (b) 18.00 mass
(c) 18.10 mass (d) 19.00 mass

Show your work.

$$(4.00)(0.40) + (20.10)(0.40) =$$

For questions 14 - 17, use the following key:
Each answer may be used once, more than once,
or not at all.

(a) alpha
(b) beta
(c) gamma
(d) alpha and beta, but not gamma

(e) a high energy beam of light
(f) free electrons & free protons
(g) a high speed electron
(h) used by Ernest Rutherford as a "probe"

For questions 18 - 21, use the following key:
Each answer may be used once, more than once,
or not at all.

(a) John Dalton
(b) Ernest Rutherford
(c) J.J. Thomson
(d) Democritus

(e) a model of the atom has been called the "planetary model"

(f) a model of the atom has been called the "billiard ball" model

(g) the model name is called the "atom"

(h) the philosophical idea included the name "atom"
(i) an addition to the atomic theory for atoms to have
both positive and negative parts

(j) the name of the Na_2O_2 ion

(k) sulfide (l) sulfate
(m) nitrate (n) hydrogen sulfide

14. Consider the following sentence. In the
blank, indicate below is present:
(a) The particle contains 30 protons

(b) The particle has a mass number of 30
(c) The particle has an atomic number of 30
(d) The particle contains 30 neutrons

15. Which elements did Mendeleev leave spaces for in
his periodic table?
6a 6b Sc

16. If copper metal is a mixture two isotopes, Cu-63,
mass = 62.994 g and Cu-65, mass = 64.924 g.
What is the mass of copper in 60.000 grams?

Calculate the "percentages" of the two isotopes of
copper. Show your work.

17. What is the formula of the ionic compound formed
between Mg and Br?
(a) MgBr (b) MgBr₂
(c) MgBr₃ (d) Mg₂Br₃

18. What is the formula of the ionic compound formed
between Ca and F?
(a) CaF₂ (b) CaF₃
(c) CaF₄ (d) CaF₅

19. What is the name of the Na_2O_2 ion?
(a) sulfide (b) sulfate
(c) nitrate (d) hydrogen sulfide

25.3
Percent

$$2.5(62.925)(x) + (49.928)(1-x) = 64.546$$

$$x = 0.1911$$

$$\text{Cu-63} \rightarrow 19.11\%$$

$$\text{Cu-65} \rightarrow 80.89\%$$

Ap Bio Pogil Chapter 48 Answer Key

M Carnoy

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