

Chapter 2 Review

1. Correct properties are characteristics of metals.

Which property means that you can pound the substance into a ball?

- a) ductility b) malleability
c) conductivity d) luster

2. Which of the following is a period?

- a) 2 b) 10 c) 5 d) 18 e) 16

3. Which of the following is a transition metal?

- a) Cl b) Fe c) P d) Ca e) C

4. Which of the following is an alkali metal?

- a) Mg b) Sr c) Rb d) Al e) Br

5. Which of the following is an lanthanide?

- a) Sr b) Ba c) Cs d) P e) W

6. Which element has the highest melting point?

- a) Pb b) Fe c) Cu d) Hg e) Mg

7. Which ions exist in the

- a) negative electrode b) positive electrode

- c) power source d) gas inside the tube

8. In a cathode ray tube, electrons are being emitted

- a) from a positively charged plate

- b) from a negatively charged plate

9. Listed below are the charges and masses of four particles. Which one will be deflected the least in a mass spectrometer?

- a) +0.2 amu b) +1.1 amu

- c) +4.4 amu d) +1.1 amu

10. In a million oil drop type experiments, the charge

on that oil droplets (Coulombs) was found to be

- 1.60 $\times 10^{-19}$ Coulombs 1.60 $\times 10^{-19}$ Coulombs
3.20 $\times 10^{-19}$ Coulombs 3.20 $\times 10^{-19}$ Coulombs
4.80 $\times 10^{-19}$ Coulombs 4.80 $\times 10^{-19}$ Coulombs
6.40 $\times 10^{-19}$ Coulombs 6.40 $\times 10^{-19}$ Coulombs

What is the charge on the electron according to this experiment?

- a) 1.1×10^{-19} Coulombs b) 1.60×10^{-19} Coulombs

- c) 3.20×10^{-19} Coulombs d) 6.40×10^{-19} Coulombs

11. Pictured below is a column of the Rutherford experiment. Which scattered particles gives the

best evidence for the nuclear atom?



- a) a b) b c) c d) d e) e

12. Which of the following is an isotope of the element

with 28 protons (p=28) and 32 neutrons (n=32)?

- a) element 22 b) element 32

- c) element 60 d) element 68

13. The imaginary element X has the following natural

abundances and isotopic masses. What is the atomic

mass of X?

- a) 15.02 amu b) 15.03 amu

- c) 15.04 amu d) 15.05 amu

Show your work.

$$(4.02 \times 0.40) + (26.10 \times 0.60) =$$

$$25.3 \text{ amu}$$

For questions 14 - 17, use the following key:

(each answer may be used once, more than once,

or not at all)

- a) alpha
b) beta
c) gamma
d) alpha and beta, but not gamma

14. A high energy form of light

15. Two protons & two neutrons

16. A high speed electron

17. Used by Ernest Rutherford as a "probe"

For questions 18 - 22, use the following key:

(each answer may be used once, more than once,

or not at all)

- a) John Dalton
b) Ernest Rutherford
c) J.J. Thomson
d) Democritus

18. The model of the atom has been called the "plum

pudding" model

19. The model of the atom has been called the "billiard

ball" model

20. The smallest matter is actually no matter

21. The philosophical idea included the term "atomos"

22. He added to the atomic theory the idea that atoms

had positive and negative parts

 23. Consider the following notation: $^{63}_{29}\text{Cu}$

Which statement below is correct?

- a) The particle contains 63 protons
b) The particle has a mass number of 36
c) The particle has an atomic number of 29
d) The particle contains 29 neutrons

24. Which elements did Mendeleev leave spaces for in

his periodic table?

 25. Copper metal is a mixture two isotopes, $^{63}_{29}\text{Cu}$,

 mass = 62.9296 u and $^{65}_{29}\text{Cu}$, mass = 64.9278 u

The natural mass of copper is 63.546 g/mol

Calculate the % abundance of the two isotopes of

copper. Show your work.

26. What is the formula of the ionic compound formed

between Mg and Br?

- a) MgBr b) MgBr₂

- c) Mg₂Br d) Mg₂Br₂

27. What is the formula of the ionic compound formed

between Ca and P?

- a) CaP₂ b) Ca₃P₂

- c) CaP d) Ca₃P₄

 28. What is the name of the Mg^{2+} ion?

- a) sulfur b) sulfur anion

- c) sulfur d) hydrogen sulfur

- e) sulfur f) sulfur

$$25) (62.9296)(x) + (64.9278)(1-x) = 63.546$$

$$x = 0.1911$$

$$\text{Cu-63} \rightarrow 19.11\%$$

$$\text{Cu-65} \rightarrow 80.89\%$$

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M Carnoy



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