

Analyzing Oxidation-Reduction Reactions

When chemical bonds form, electrons are either lost, gained or shared. Metals lose electrons. This is what happens when iron rusts. When the iron, a metal, combines with oxygen, a non metal, to form rust, it loses electrons. This process is called oxidation even when the nonmetal is not oxygen. Nonmetals gain electrons causing their oxidation states to go down. This is called reduction. It is possible to tell what was oxidized and what was reduced in a chemical reaction by checking the oxidation states of the elements before and after the reaction. The element that has an increase in oxidation state was oxidized while the one that has a decrease in oxidation state was reduced.



Example:



For each of the examples below, determine the oxidation states of the elements on both sides of the equation. Then determine which element was oxidized and which was reduced. Write your answer in the space provided.

Reaction	Element:	
	Oxidized	Reduced
Example: $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$ $\overset{0}{\text{Cu}} + 2\overset{+1}{\text{Ag}}\overset{+5}{\text{N}}\overset{-2}{\text{O}_3} \rightarrow \overset{+2}{\text{Cu}}(\overset{+5}{\text{N}}\overset{-2}{\text{O}_3})_2 + 2\overset{0}{\text{Ag}}$	Cu	Ag
1. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$		
2. $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$		
3. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$		
4. $2\text{K}_2\text{Cr}_2\text{O}_7 + 2\text{H}_2\text{O} + 18\text{H}^+ \rightarrow 4\text{H}^+\text{OH} + 2\text{Cr}_2\text{O}_3 + 3\text{SO}_2$		

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Raymond Chang



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CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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Holt Chemistry Ralph Thomas Myers,2004

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A System of Instruction in Quantitative Chemical Analysis Karl Remigius Fresenius,1882

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The Students' Guide in Quantitative Analysis Henry Carrington Bolton,1881

A Systematic Handbook of Volumetric Analysis Francis Sutton,1890

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Review of Metal Literature American Society for Metals,1966

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Journal of Chemical Education ,1927

Textbook of Macro and Semimicro Qualitative Inorganic Analysis Arthur Israel Vogel,G. Svehla,1987

Originally written in 1937 this edition contains several new features including marginal notes highlighting facts and summarizing reactions extended chapters on preliminary tests dissolution and fusion of samples and updated nomenclature

Metallurgical Analysis Nathaniel Wright Lord,Dana James Demorest,1924

Flavor Analysis Cynthia J. Mussinan,Michael J. Morello,1998 Filled with practical examples this volume illustrates innovative flavor analysis techniques used by today s leaders in food chemistry It covers flavor analysis for apples beef citrus coffee cheese extruded pet foods fungi milk and Maillard reaction systems and the researchers come from throughout the industrialized world Eleven of the chapters illustrate techniques for isolating volatile compounds from complex food matrices including micro scale liquid liquid extraction headspace sampling solid phase microextraction supercritical fluid extraction and thermo desorption The chapters devoted to analytical characterization include analyses of Amadori compounds sulfur compounds chiral compounds

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