

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

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Luann Marie Decker



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Table of Contents 12 Stoichiometry Practice Problem Worksheetanswer Key

1. Understanding the eBook 12 Stoichiometry Practice Problem Worksheetanswer Key
 - The Rise of Digital Reading 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12 Stoichiometry Practice Problem Worksheetanswer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Personalized Recommendations
 - 12 Stoichiometry Practice Problem Worksheetanswer Key User Reviews and Ratings
 - 12 Stoichiometry Practice Problem Worksheetanswer Key and Bestseller Lists
5. Accessing 12 Stoichiometry Practice Problem Worksheetanswer Key Free and Paid eBooks
 - 12 Stoichiometry Practice Problem Worksheetanswer Key Public Domain eBooks
 - 12 Stoichiometry Practice Problem Worksheetanswer Key eBook Subscription Services
 - 12 Stoichiometry Practice Problem Worksheetanswer Key Budget-Friendly Options
6. Navigating 12 Stoichiometry Practice Problem Worksheetanswer Key eBook Formats

- ePub, PDF, MOBI, and More
 - 12 Stoichiometry Practice Problem Worksheetanswer Key Compatibility with Devices
 - 12 Stoichiometry Practice Problem Worksheetanswer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Highlighting and Note-Taking 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Interactive Elements 12 Stoichiometry Practice Problem Worksheetanswer Key
 8. Staying Engaged with 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12 Stoichiometry Practice Problem Worksheetanswer Key
 9. Balancing eBooks and Physical Books 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12 Stoichiometry Practice Problem Worksheetanswer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Setting Reading Goals 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Fact-Checking eBook Content of 12 Stoichiometry Practice Problem Worksheetanswer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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