

AQA Required Practical Summaries - Chemistry Paper 1

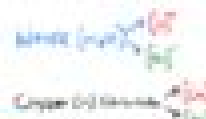
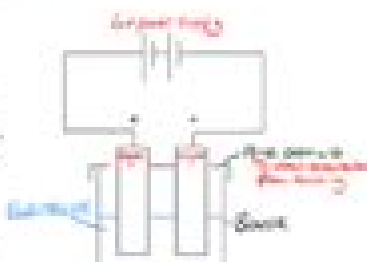
RP3: Electrolysis

Aims Investigate what happens when two different aqueous solutions are electrolysed using inert (unreactive) electrodes



Method

1. Pour 50 cm³ of copper (II) chloride solution into a beaker.
2. Insert carbon rods ensuring they do not touch.
3. Attach crocodile leads to the rods and connect them a power supply (4 V).
4. Switch on the power supply and initial observations.
5. Use forceps to hold blue litmus (anode) to identify the elements.
6. Record all observations in a table.
7. Rinse the apparatus and collect.
8. Repeat steps 1-7 using 50 cm³.



When ionic substances dissolve in water, they split into ions, as does the water itself, allowing ions to move freely and carry charge. During current flow, ions migrate to oppositely charged electrodes.



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RP 3: Electrolysis - Common Questions

You may be asked to identify products formed during electrolysis

This table (also shows) information about the products of the electrolysis of an aqueous solution of copper chloride.

	Product at positive electrode	Product at negative electrode
Name of product	Oxygen	Copper
State of product	Gas	Solid

Complete the table above.

(5)

You may be asked what happens in the electrolyte.

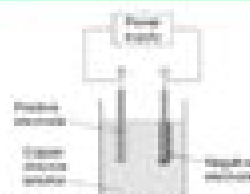
Electrolyte is any suitable where an ionic compound is either in an aqueous solution.

Explain why.

You should refer to ions and charge in your answer.

(6)

You may be asked to describe what would be observed at the electrodes.



Describe what is seen at each electrode during the electrolysis of copper chloride solution.

Positive electrode _____

Negative electrode _____

(2)

Examiners may link to qualitative calculations.

The copper chloride solution used in the investigation contained 200 grams per dm³ of solid CuCl₂ dissolved in 1 dm³ of water.

The students used 50 cm³ of copper chloride solution in each experiment.

Calculate the mass of solid copper chloride used in each experiment.

Mass = _____ g

(3)



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Sandra Mitchell, Charles Golabek

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