

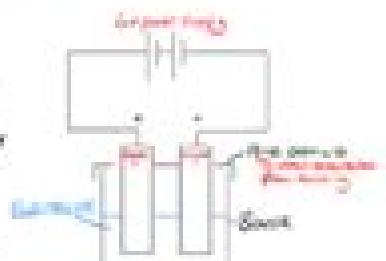
# AQA Required Practical Summaries - Chemistry Paper 1

## RP3: Electrolysis

### Method

1. Pour 50 cm<sup>3</sup> of copper (II) chloride solution into a beaker.
2. Insert carbon rods ensuring they do not touch.
3. Attach crocodile leads to the rods and connect them to a power supply (4 V).
4. Switch on the power supply and initial observations.
5. Use forceps to hold blue litmus (anode) to identify the element.
6. Record all observations in a table.
7. Rinse the apparatus and collect.
8. Repeat steps 1-7 using 50 cm<sup>3</sup>

Aim: Investigate what happens when two different aqueous solutions are electrolysed using inert (unreactive) electrodes



When ionic substances dissolve in water, they split into ions, as does the water itself, allowing ions to move freely and carry charge. During current flow, ions migrate to oppositely charged electrodes

### RP 3: Electrolysis - Common Questions

You may be asked to identify products formed during electrolysis

The table below gives information about the products of the electrolysis of an aqueous solution of copper(II) chloride.

	Product at positive electrode	Product at negative electrode
Name of product	Chlorine	
Name of product		Hydrogen

Complete the table above.

You may be asked what happens in the electrolyte.

Explain in one sentence what an ionic compound is made of in aqueous solution.

Answer:   
\_\_\_\_\_

You should refer to ions and charge in your answer.

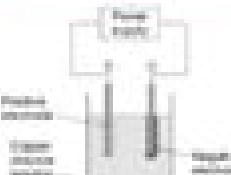
\_\_\_\_\_

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\_\_\_\_\_

You may be asked to describe what would be observed at the electrodes.



Describe what is seen at each electrode using the electrolysis of copper(II) chloride.

Anode:   
\_\_\_\_\_

Cathode:   
\_\_\_\_\_

Positive electrode:   
\_\_\_\_\_

Examiners may link to qualitative calculations.

The copper(II) chloride solution used in the investigation contained 300 grams per dm<sup>3</sup> of solid CuCl<sub>2</sub> dissolved in 1 dm<sup>3</sup> of water.

The students used 50 cm<sup>3</sup> of copper(II) chloride solution in each experiment.

Calculate the mass of solid copper(II) chloride used in each experiment.

Mass =   
\_\_\_\_\_



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**Sandra Mitchell,Charles Golabek**

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