

Name Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 , is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

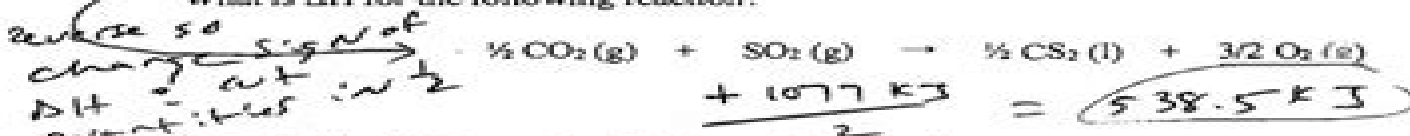
$$\frac{99.2 \text{ mol } H_2}{2 \text{ mol } H_2} \times \frac{-484 \text{ kJ}}{1} = -24,006 \text{ kJ}$$

$-24,006 \text{ kJ}$
 ~~$-24,006 \text{ kJ}$~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

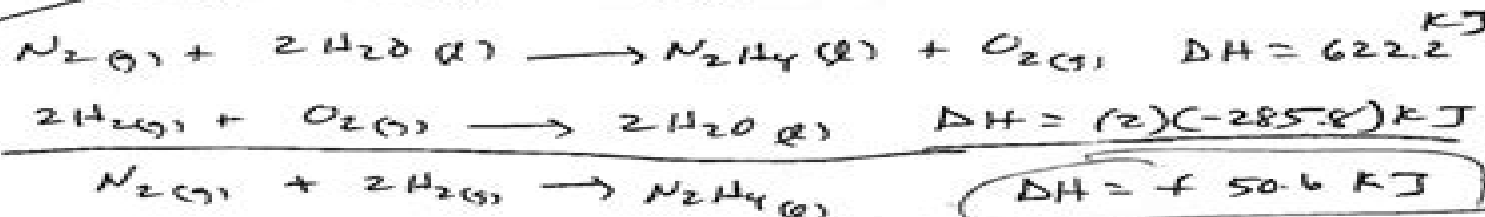
Exothermic

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

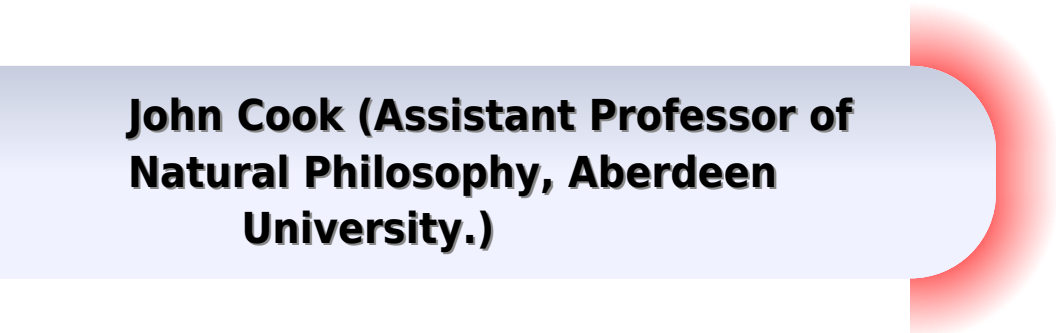
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5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



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**John Cook (Assistant Professor of
Natural Philosophy, Aberdeen
University.)**



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1. Understanding the eBook 2 Practice Problems Chemistry Answers
 - The Rise of Digital Reading 2 Practice Problems Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 2 Practice Problems Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2 Practice Problems Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2 Practice Problems Chemistry Answers
 - Personalized Recommendations
 - 2 Practice Problems Chemistry Answers User Reviews and Ratings
 - 2 Practice Problems Chemistry Answers and Bestseller Lists
5. Accessing 2 Practice Problems Chemistry Answers Free and Paid eBooks

- 2 Practice Problems Chemistry Answers Public Domain eBooks
- 2 Practice Problems Chemistry Answers eBook Subscription Services
- 2 Practice Problems Chemistry Answers Budget-Friendly Options
- 6. Navigating 2 Practice Problems Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 2 Practice Problems Chemistry Answers Compatibility with Devices
 - 2 Practice Problems Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2 Practice Problems Chemistry Answers
 - Highlighting and Note-Taking 2 Practice Problems Chemistry Answers
 - Interactive Elements 2 Practice Problems Chemistry Answers
- 8. Staying Engaged with 2 Practice Problems Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2 Practice Problems Chemistry Answers
- 9. Balancing eBooks and Physical Books 2 Practice Problems Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2 Practice Problems Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 2 Practice Problems Chemistry Answers
 - Setting Reading Goals 2 Practice Problems Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 2 Practice Problems Chemistry Answers
 - Fact-Checking eBook Content of 2 Practice Problems Chemistry Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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