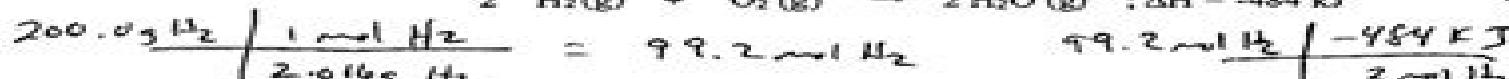


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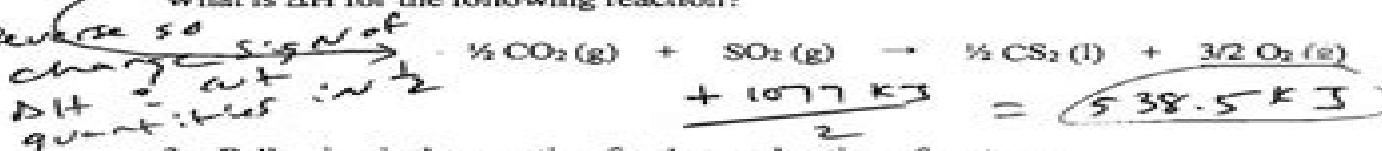
1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$  is burned?



2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



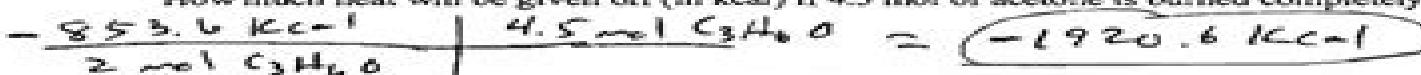
What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?



4. The specific heat of iron is 0.110 cal / g  $^{\circ}\text{C}$ . How much heat must be added to 5.2 g of iron at 25  $^{\circ}\text{C}$  to raise its temperature to 225  $^{\circ}\text{C}$ ?

$$\text{heat} = q = m \cdot c \cdot \Delta T = (5.2)(0.110 \frac{\text{cal}}{\text{g} \cdot \text{C}})(225 - 25) = 114.4 \text{ cal}$$

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



Reverse



double  
quantities



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**John Cook (Assistant Professor of  
Natural Philosophy, Aberdeen  
University.)**

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