

**19.2****HYDROGEN IONS AND ACIDITY****Section Review****Objectives**

- Classify a solution as neutral, acidic, or basic, given the hydrogen-ion or hydroxide-ion concentration
- Convert hydrogen-ion concentrations into values of pH and hydroxide-ion concentrations into values of pOH
- Describe the purpose of pH indicators

**Vocabulary**

- self-ionization
- ion-product constant for water ( $K_w$ )
- acidic solution
- basic solution
- alkaline solutions
- pH

**Key Equations**

- $K_w = [\text{H}^+] \times [\text{OH}^-] = 1.0 \times 10^{-14}\text{M}^2$
- $\text{pH} = -\log [\text{H}^+]$
- $\text{pOH} = -\log [\text{OH}^-]$
- $\text{pH} + \text{pOH} = 14$

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water molecules can \_\_\_\_1\_\_\_\_ to form hydrogen ions ( $\text{H}^+$ ) and hydroxide ions ( $\text{OH}^-$ ). The concentrations of these ions in pure water at 25°C are both equal to \_\_\_\_2\_\_\_\_ mol/L.

The pH scale, which has a range from \_\_\_\_3\_\_\_\_, is used to denote the \_\_\_\_4\_\_\_\_ concentration of a solution. On this scale, 0 is strongly \_\_\_\_5\_\_\_\_, 14 is strongly \_\_\_\_6\_\_\_\_, and 7 is \_\_\_\_7\_\_\_\_. Pure water at 25°C has a pH of \_\_\_\_8\_\_\_\_.

The \_\_\_\_9\_\_\_\_ constant for water has a value of  $1.0 \times 10^{-14}$ . Thus, the product of the concentrations of \_\_\_\_10\_\_\_\_ ions and \_\_\_\_11\_\_\_\_ ions in aqueous solution will always equal  $1.0 \times 10^{-14}$ .

1. ionize
2.  $1 \times 10^{-7}$  mol/L
3. 0-14
4.  $\text{H}^+$
5. acidic
6. basic
7. neutral
8. 7
9. ion product
10.  $\text{H}^+ / \text{H}_3\text{O}^+$
11.  $\text{OH}^-$

## 19 3 Section Review Answers

**Peter R. Mansoor**



### 19 3 Section Review Answers:

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