

19.2**HYDROGEN IONS AND ACIDITY****Section Review****Objectives**

- Classify a solution as neutral, acidic, or basic, given the hydrogen-ion or hydroxide-ion concentration
- Convert hydrogen-ion concentrations into values of pH and hydroxide-ion concentrations into values of pOH
- Describe the purpose of pH indicators

Vocabulary

• self-ionization	• ion-product constant for water (K_w)	• basic solution
• neutral solution	• acidic solution	• alkaline solutions
		• pH

Key Equations

- $K_w = [H^+] \times [OH^-] = 1.0 \times 10^{-14} M^2$
- $pOH = -\log [OH^-]$
- $pH = -\log [H^+]$
- $pOH = -\log [OH^-]$
- $pH + pOH = 14$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water molecules can 1 to form hydrogen ions (H^+) and hydroxide ions (OH^-). The concentrations of these ions in pure water at 25°C are both equal to 2 mol/L.

The pH scale, which has a range from 3, is used to denote the 4 concentration of a solution. On this scale, 0 is strongly 5, 14 is strongly 6, and 7 is 7. Pure water at 25°C has a pH of 8.

The 9 constant for water has a value of 1.0×10^{-14} . Thus, the product of the concentrations of 10 ions and 11 ions in aqueous solution will always equal 1.0×10^{-14} .

1. ionize
2. 1×10^{-7} mol/L
3. 0-14
4. H^+
5. acidic
6. basic
7. neutral
8. 7
9. ion product
10. H^+ / OH^-
11. OH^-

19 3 Section Review Answers

Peter R. Mansoor

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