

## 10-3 Practice Problems

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- Find the percentage composition of a compound that contains 1.94-g of carbon, 0.48-g of hydrogen, and 2.58-g of sulfur in a 5.00-g sample of the compound.
- A sample of an unknown compound with a mass of 0.847-g has the following composition: 50.51 percent fluorine and 49.49 percent iron. When this compound is decomposed into its elements, what mass of each element would be recovered?
- Find the percentage composition of a compound that contains 2.63-g of carbon, 0.370-g of hydrogen, and 0.580-g of oxygen in a 3.58-g sample of the compound.
- A sample of an unknown compound with a mass of 2.876-g has the following composition: 66.07 percent carbon, 6.71 percent hydrogen, 4.06 percent nitrogen, and 23.16 percent oxygen. What is the mass of each element in this compound?
- Find the percentage composition of a compound that contains 2.7368-g of chlorine, 0.4116-g of oxygen, and 0.7971-g of phosphorus in a 3.9460-g sample of the compound.
- Find the percentage composition of a compound that contains 1.51-g of chromium, 1.13-g of potassium, and 1.82-g of oxygen in a 4.26-g sample of the compound.
- A sample of a compound that has a mass of 0.632-g is analyzed. The sample is found to be made up of oxygen and fluorine only. Given that the sample contains 0.128-g of oxygen, calculate the percentage composition of the compound.
- What is the percentage composition of a carbon-oxygen compound, given that a 95.2-g sample of the compound contains 40.8-g of carbon and 54.4-g of oxygen?
- What is the percentage composition of a sulfur-chlorine compound, given that a 30.9-g sample of the compound contains 9.63-g of sulfur and 21.3-g of chlorine?
- Determine the empirical formula of a compound containing 2.644-g of gold and 0.476-g of chlorine.
- Determine the empirical formula of a compound containing 0.928-g of gallium and 0.612-g of phosphorus.
- Determine the empirical formula of a compound containing 1.723-g of carbon, 0.289-g of hydrogen, and 0.459-g of oxygen.

# 19 3 Practice Problems Chemistry Answers

**S Ashworth**



### 19 3 Practice Problems Chemistry Answers:

*Survival Guide to General Chemistry* Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

**Student Solutions Manual** Darrell D. Ebbing, 2002-07 Provides solutions to exercises solutions to odd numbered practice problems general problems and cumulative skills problems plus answers to review questions Ebbing Darrell D. Ebbing, 1993 **Partial Solutions Manual** Darrell D. Ebbing, David Bookin, 1998

*Basic Mathematics* Howard A. Silver, 1995-08 Chemistry Workbook For Dummies Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence *Holt Science and Technology 2002* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2002 **Clinical Chemistry** Michael L. Bishop, Janet L. Duben-Engelkirk, Edward P. Fody, 2000 Written in a concise readable style the Fourth Edition of this leading text continues to set the standard in the constantly evolving field of clinical chemistry Completely revised and updated this text reflects the latest developments in clinical chemistry Recent advances in quality assurance PCR and laboratory automation receive full coverage The

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*National Library of Medicine Current Catalog* National Library of Medicine (U.S.), 1971 First multi year cumulation covers six years 1965-70      *The American Catalogue*, 1880 American national trade bibliography      *The American Bookseller*, 1886      *The Journal of Education*, 1893

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