

REF ID:

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

Chapters 17 & 18

Classification of the Standards

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)

9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)

9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)

9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)

9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)

9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CSS: 8c)

9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9a)

Wesleyan

- Condensation
 - Freezing
 - Vaporization
 - Melting
 - Boiling
 - Sublimation
 - Deposition
 - Heat
 - Temperature
 - Thermochemistry
 - Exothermic
 - Endothermic
 - Calorimeter
 - Joule
 - Specific heat
 - $q = m \cdot c \cdot \Delta T$
 - $q = m \cdot \Delta H_{\text{exothermic}}$
 - $q = m \cdot \Delta H_{\text{endothermic}}$
 - Heat of reaction
 - Le Chatelier's principle
 - Equilibrium
 - Activation Energy
 - Catalyst

Study Guide (use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?
Must collide with each other. Must collide with enough energy. & in the correct orientation

2. What is the importance of the activation energy in a chemical reaction?

Amount of energy needed for a reaction to occur

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction.
 $2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$

- a. Increase temp: Increase, molecules collide more
 - b. Increase HCl: Increase, more collisions because more reactants to collide
 - c. Increase pressure: Hard to say, wouldn't affect the reactants directly because none of them are gases. One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow down. Tough call. (this one won't be on the test)
 - d. Remove $MgCl_2$: Increase, less products means less chance for them to interfere with reactant collisions
 - e. Add H_2 : Decrease, more products mean they interfere more with reactants and the reactants would

One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough cell.. (this one won't be on the test).

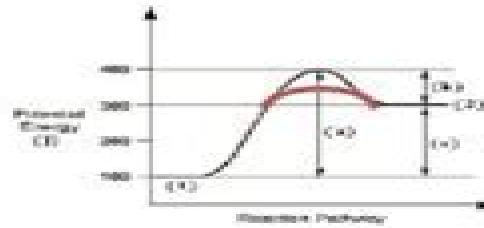
- d. Remove $MgCl_2$: Increase, less products means less chance for them to interfere with reactant collisions
e. Add H_2 : Decrease, more products mean they interfere more with reactants and the reactants would collide less.

4. Identify the following terms with their correct letter from the graphs:

- Reactants- 1
 - Activation energy- A
 - Is this reaction exothermic or endothermic?
 - Products- 2

Endothermic

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph).



18 Rates Of Reaction Section Answer Key

**George A. Williams, Richard
Barnes, Ray Doerhoff, Max Bolen**

18 Rates Of Reaction Section Answer Key:

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