

## Section Review

## Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium
- Identify three stresses that can change the equilibrium position of a chemical system
- Explain what the value of  $K_{eq}$  indicates about the position of equilibrium

## Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Châtelier's principle
- equilibrium constant ( $K_{eq}$ )

## Key Equation

$$K_{eq} = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

When  $aA + bB \rightleftharpoons cC + dD$

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

1. A reversible reaction is one in which the reaction can proceed in both directions. The reaction is reversible if the reaction is reversible.

2. The equilibrium constant,  $K_{eq}$ , is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its coefficient in the balanced chemical equation. The value of  $K_{eq}$  is constant for a given reaction at a given temperature.

1. reversible
2. products
3. reactants
4. products
5. reactants
6. products
7. reactants
8. products
9. reactants
10. products
11. reactants
12. products
13. reactants
14. products
15. reactants
16. products
17. reactants
18. products
19. reactants
20. products

# 18 3 Reversible Reactions And Equilibrium Answer Key

**John McMurry, Robert C. Fay**



### 18 3 Reversible Reactions And Equilibrium Answer Key:

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