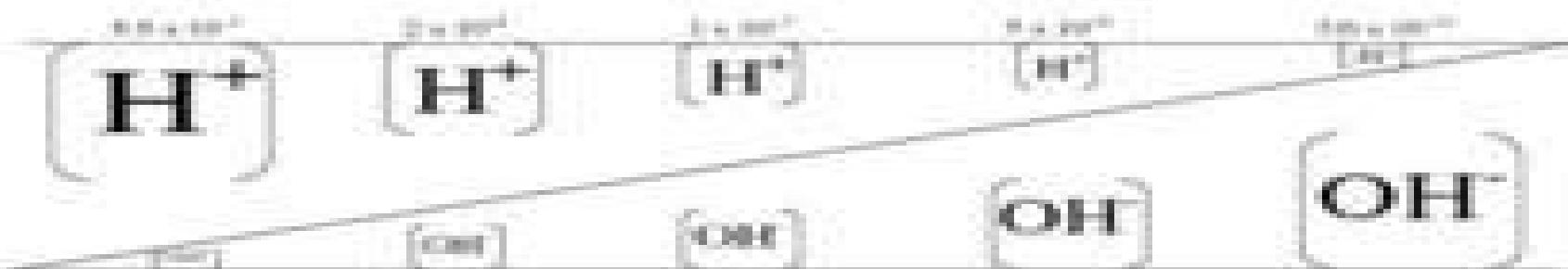


1.7 • Acid-Base Equilibrium

pH = -log₁₀[H⁺]

- What is the pH of each solution given the following $\text{pH} = -\log_{10}[\text{H}^+]$?
 - Calculated pH: _____
 - Actual pH: _____
- Knowing that $\text{pH} = -\log_{10}[\text{H}^+]$ where $[\text{H}^+] = 10^{-7} \text{ M}$, calculate pH for each of the following concentrations for the H^+ ion. $\text{pH} = -\log_{10}[\text{H}^+]$ is called the pH because it measures the concentration of hydronium ions. $\text{pH} = -\log_{10}[\text{H}_3\text{O}^+]$ is called the pH because it measures the concentration of hydronium ions.
- Given the $\text{pH} = 7.00$, calculate $[\text{H}^+]$ _____.
- $\text{pH} = -\log_{10}[\text{H}^+]$ requires _____.
- Given the $\text{pH} = 10.00$, calculate $[\text{H}^+]$ _____.
 - Calculate the pH _____.
 - Calculate the pH _____.
 - Calculate the pH _____.
- Given the $\text{pH} = 10.00$, calculate $[\text{H}^+]$ _____.
 - Calculate the pH _____.
 - Calculate the pH _____.
 - Calculate the pH _____.
- Calculate the pH for the following solutions:
0.001 M H_2O : $\text{pH} =$ _____
0.0001 M H_2O : $\text{pH} =$ _____
0.00001 M H_2O : $\text{pH} =$ _____
0.000001 M H_2O : $\text{pH} =$ _____
0.0000001 M H_2O : $\text{pH} =$ _____
0.00000001 M H_2O : $\text{pH} =$ _____
0.000000001 M H_2O : $\text{pH} =$ _____
0.0000000001 M H_2O : $\text{pH} =$ _____
0.00000000001 M H_2O : $\text{pH} =$ _____

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Wolfgang Guggemos

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