Name _____ Date _____ Class _____

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

# Solutions

## Section 15.1 What are solutions?

*In your textbook, read about the characteristics of solutions.*

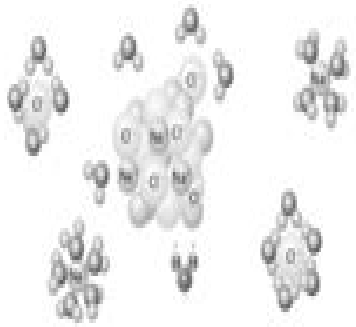
**Use each of the terms below just once to complete the passage.**

| | | | |
|---|---|---|---|
| immiscible | liquid | soluble | solution |
| insoluble | miscible | solute | solvent |

Air is an (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute, and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an (4) soluble substance. A substance that does not dissolve is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

*Read about solvation in aqueous solutions in your textbook.*

**The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.**



9. Hydration is solvation in which the solvent is water. What is solvation?

Solvation is the process of surrounding solute particles with solvent particles to form a solution.



Name _____ Date _____ Class _____



**Section 15.1** *continued*

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the orientation of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

*In your textbook, read about heat of solution, solubility, and factors that affect solubility.*

**For each statement below, write *true* or *false*.**

true 14. The overall energy change that occurs when a solution forms is called the heat of solution.

false 15. Solubility is a measure of the minimum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

false 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

true 17. In a saturated solution, solvation and crystallization are in equilibrium.

true 18. Additional solute can be dissolved in an unsaturated solution.

true 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.



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**Table of Contents 15 Study Guide For Content Mastery Solutions**


1. Understanding the eBook 15 Study Guide For Content Mastery Solutions
    - The Rise of Digital Reading 15 Study Guide For Content Mastery Solutions
    - Advantages of eBooks Over Traditional Books
2. Identifying 15 Study Guide For Content Mastery Solutions
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an 15 Study Guide For Content Mastery Solutions
    - User-Friendly Interface
4. Exploring eBook Recommendations from 15 Study Guide For Content Mastery Solutions
    - Personalized Recommendations
    - 15 Study Guide For Content Mastery Solutions User Reviews and Ratings
    - 15 Study Guide For Content Mastery Solutions and Bestseller Lists
5. Accessing 15 Study Guide For Content Mastery Solutions Free and Paid eBooks
    - 15 Study Guide For Content Mastery Solutions Public Domain eBooks
    - 15 Study Guide For Content Mastery Solutions eBook Subscription Services

- 15 Study Guide For Content Mastery Solutions Budget-Friendly Options
6. Navigating 15 Study Guide For Content Mastery Solutions eBook Formats
   - ePub, PDF, MOBI, and More
   - 15 Study Guide For Content Mastery Solutions Compatibility with Devices
   - 15 Study Guide For Content Mastery Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
   - Adjustable Fonts and Text Sizes of 15 Study Guide For Content Mastery Solutions
   - Highlighting and Note-Taking 15 Study Guide For Content Mastery Solutions
   - Interactive Elements 15 Study Guide For Content Mastery Solutions
8. Staying Engaged with 15 Study Guide For Content Mastery Solutions
   - Joining Online Reading Communities
   - Participating in Virtual Book Clubs
   - Following Authors and Publishers 15 Study Guide For Content Mastery Solutions
9. Balancing eBooks and Physical Books 15 Study Guide For Content Mastery Solutions
   - Benefits of a Digital Library
   - Creating a Diverse Reading Collection 15 Study Guide For Content Mastery Solutions
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine 15 Study Guide For Content Mastery Solutions
    - Setting Reading Goals 15 Study Guide For Content Mastery Solutions
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 15 Study Guide For Content Mastery Solutions
    - Fact-Checking eBook Content of 15 Study Guide For Content Mastery Solutions
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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