

**15.2****HOMOGENEOUS AQUEOUS SYSTEMS****Section Review****Objectives**

- Distinguish between a solvent and a solute
- Describe what happens in the solution process
- Explain why all ionic compounds are electrolytes
- Demonstrate how to write the formula for a hydrate

**Vocabulary**

• aqueous solution	• solvation	• strong electrolyte
• solvent	• electrolyte	• weak electrolyte
• solute	• nonelectrolyte	• hydrate

**Key Equation**

$$\bullet \text{ Percent H}_2\text{O} = \frac{\text{mass of water}}{\text{mass of hydrate}} \times 100\%$$

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water is a polar liquid and an excellent 1 for many substances. Aqueous solutions are 2 mixtures of ions or molecules in water. The solubility of a solute depends on solute-solvent interactions. A good rule to remember is 3.

Substances that dissolve as ions are known as 4. A solute that is completely ionized in solution is a 5 electrolyte. A weak electrolyte is only 6 ionized. A solution of an electrolyte will 7 an electric current, whereas a solution of a 8 is nonconducting.

Many crystals are 9; they contain water of hydration. In the process called 10, the water of hydration is lost from a hydrate that is exposed to air.

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## 15 2 Homogeneous Aqueous Systems Answer Key

**RC Schank**

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## **Table of Contents 15 2 Homogeneous Aqueous Systems Answer Key**

1. Understanding the eBook 15 2 Homogeneous Aqueous Systems Answer Key
  - The Rise of Digital Reading 15 2 Homogeneous Aqueous Systems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 15 2 Homogeneous Aqueous Systems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 15 2 Homogeneous Aqueous Systems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 15 2 Homogeneous Aqueous Systems Answer Key
  - Personalized Recommendations
  - 15 2 Homogeneous Aqueous Systems Answer Key User Reviews and Ratings
  - 15 2 Homogeneous Aqueous Systems Answer Key and Bestseller Lists
5. Accessing 15 2 Homogeneous Aqueous Systems Answer Key Free and Paid eBooks
  - 15 2 Homogeneous Aqueous Systems Answer Key Public Domain eBooks
  - 15 2 Homogeneous Aqueous Systems Answer Key eBook Subscription Services
  - 15 2 Homogeneous Aqueous Systems Answer Key Budget-Friendly Options
6. Navigating 15 2 Homogeneous Aqueous Systems Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- 15 2 Homogeneous Aqueous Systems Answer Key Compatibility with Devices
- 15 2 Homogeneous Aqueous Systems Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 15 2 Homogeneous Aqueous Systems Answer Key
- Highlighting and Note-Taking 15 2 Homogeneous Aqueous Systems Answer Key
- Interactive Elements 15 2 Homogeneous Aqueous Systems Answer Key

8. Staying Engaged with 15 2 Homogeneous Aqueous Systems Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 15 2 Homogeneous Aqueous Systems Answer Key

9. Balancing eBooks and Physical Books 15 2 Homogeneous Aqueous Systems Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 15 2 Homogeneous Aqueous Systems Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 15 2 Homogeneous Aqueous Systems Answer Key

- Setting Reading Goals 15 2 Homogeneous Aqueous Systems Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 15 2 Homogeneous Aqueous Systems Answer Key

- Fact-Checking eBook Content of 15 2 Homogeneous Aqueous Systems Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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