

Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent

$$14.8 \text{ g C}_3\text{H}_8 \times \frac{1 \text{ mol C}_3\text{H}_8}{44.09 \text{ g C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol CO}_2} = 1.55 \text{ g H}_2\text{O}$$

O_2 is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.84 \text{ g H}_2\text{O}$$

c) determine the number of grams of H_2O produced

1.55 g H_2O

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol C}_3\text{H}_8}{5 \text{ mol O}_2} \times \frac{44.09 \text{ g C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} = .0948 \text{ g C}_3\text{H}_8 \text{ used}$$

$$14.8 \text{ g C}_3\text{H}_8 - 0.948 \text{ g C}_3\text{H}_8 = 13.9 \text{ g C}_3\text{H}_8 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{294.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g Na}_2\text{SO}_4 \times \frac{1 \text{ mol Na}_2\text{SO}_4}{126.05 \text{ g Na}_2\text{SO}_4} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol NaOH}} = 15.8 \text{ g Na}_2\text{SO}_4$$

14 Limiting Reagents Answers

**Christopher R. Hren, John T.
Moore, Peter J. Mikulecky**

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Mastering Chemistry Gerald W. Gibson, 1975

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Table of Contents 14 Limiting Reagents Answers

1. Understanding the eBook 14 Limiting Reagents Answers
 - The Rise of Digital Reading 14 Limiting Reagents Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 14 Limiting Reagents Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 14 Limiting Reagents Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 14 Limiting Reagents Answers
 - Personalized Recommendations
 - 14 Limiting Reagents Answers User Reviews and Ratings
 - 14 Limiting Reagents Answers and Bestseller Lists
5. Accessing 14 Limiting Reagents Answers Free and Paid eBooks
 - 14 Limiting Reagents Answers Public Domain eBooks
 - 14 Limiting Reagents Answers eBook Subscription Services
 - 14 Limiting Reagents Answers Budget-Friendly Options
6. Navigating 14 Limiting Reagents Answers eBook Formats

- ePub, PDF, MOBI, and More
- 14 Limiting Reagents Answers Compatibility with Devices
- 14 Limiting Reagents Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 14 Limiting Reagents Answers
- Highlighting and Note-Taking 14 Limiting Reagents Answers
- Interactive Elements 14 Limiting Reagents Answers

8. Staying Engaged with 14 Limiting Reagents Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 14 Limiting Reagents Answers

9. Balancing eBooks and Physical Books 14 Limiting Reagents Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 14 Limiting Reagents Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 14 Limiting Reagents Answers

- Setting Reading Goals 14 Limiting Reagents Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 14 Limiting Reagents Answers

- Fact-Checking eBook Content of 14 Limiting Reagents Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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