

**BAGONG SILANG NATIONAL HIGH SCHOOL**  
**1<sup>ST</sup> QUARTER EXAMINATION**  
**PHYSICAL SCIENCE**

NAME: \_\_\_\_\_  
GRADE: \_\_\_\_\_

DATE: \_\_\_\_\_  
SCORE: \_\_\_\_\_

Directions: Write the letter of the correct answer on the space provided.

- \_\_\_\_\_ 1. It is also known as nuclear fusion and the formation of new nuclei actions in the early stages of development of the universe.  
a. Nucleosynthesis                      b. S-Process                      c. R-Process                      d. Singularity
- \_\_\_\_\_ 2-4. There are two categories of star, the average star and massive star. The life cycle of these two stars are common from 1<sup>st</sup> to 3<sup>rd</sup> stages namely nebula, protostar, and main sequence star. What will be the next stages in the life of an average star until it reaches its death?  
a. Red super giant → Supernova → Black hole → neutron star 0  
b. Red giant → white dwarf → black dwarf 3  
c. Red giant → white dwarf → black hole 2  
d. Red supergiant → white dwarf → black dwarf 1
- \_\_\_\_\_ 5. He synthesized element with atomic number 43 using a linear particle accelerator.  
a. Ernest Rutherford      b. Ernest Lawrence      c. Dmitri Mendeleev      d. John Dalton
- \_\_\_\_\_ 6. The following are light elements EXCEPT:  
a. helium                      b. hydrogen                      c. lithium                      d. iron
- \_\_\_\_\_ 7. Which of the following molecule is nonpolar?  
a. NaCl                      b. HCl                      c. CO<sub>2</sub>                      d. NH<sub>3</sub>
- \_\_\_\_\_ 8. What type of chemical bond holds the atoms of water molecule together?  
a. Hydrogen bond      b. Ionic bond                      c. Polar covalent bond      d. Nonpolar covalent bond
- \_\_\_\_\_ 9. Which of the following IMFAs is considered as the weakest?  
a. H-bonding                      b. Ion-dipole                      c. Dipole-dipole                      d. London forces
- \_\_\_\_\_ 10. Which of the following statements is TRUE about nonpolar molecules?  
a. high boiling point                      c. low surface tension  
b. high melting point                      d. low vapor pressure
- \_\_\_\_\_ 11. Forces that are present between an ion and a polar molecule  
a. ion-dipole                      b. dipole-dipole                      c. dispersion forces                      d. hydrogen bonding
- \_\_\_\_\_ 12. H-bonding forms when the substances involved are polar and there are \_\_\_\_\_.  
a. H-atoms of a molecule interact with to O, N, F of other molecules.  
b. C-atoms attached to O, N, F.  
c. central atoms with O, N, F as attached atoms.  
d. unshared pair of electrons in the central atom.
- \_\_\_\_\_ 13. The amount of energy required to stretch or increase the surface of a liquid by a unit area  
a. specific heat                      b. surface tension                      c. vapor pressure                      d. heat of vaporization
- \_\_\_\_\_ 14. Liquids can form spherical elastic film to minimize surface area. What intermolecular force is responsible for the formation of this film in water?  
a. H-bonding                      b. ion-induced dipole                      c. dipole-induced dipole      d. London dispersion force
- \_\_\_\_\_ 15. Which group of biomolecules carries and passes on the hereditary information of the organisms?  
a. carbohydrates                      b. lipids                      c. nucleic acids                      d. proteins
- \_\_\_\_\_ 16. What are the elements that make up carbohydrates?  
a. C and H                      b. C, H and N                      c. C, H and O                      d. C, H, O and N
- \_\_\_\_\_ 17-19. Is carbohydrates one of the most important biological macromolecules needed by a human body?  
a. Yes, because it is the primary source of energy. 3  
b. Yes, because it can be a storage of energy for a long period of time. 1  
c. Yes, because it can be found in rice, cereals, fruits and pastas. 2  
d. Yes, because it carries genetic information. 0
- \_\_\_\_\_ 20. The measure of how fast is the change in the concentration of the reactants or products in a chemical reaction.  
a. Activation Energy                      b. Activation Rate                      c. Chemical Reaction                      d. Concentration
- \_\_\_\_\_ 21. A chemical change that occurs when two or more substances combine to form a new substance.  
a. Activation Energy                      b. Chemical Kinetics                      c. Chemical Reaction                      d. Product

# Section 2review Answer Key Physical Science

**Ralph M. Feather**



## **Section 2review Answer Key Physical Science:**

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**[grade] 3, units E and F, teacher's ed** ,2000      **Glencoe Physical Science** McGraw Hill,1999 2000 2005 State Textbook Adoption      **Educational Times** ,1896      Musical Opinion and Music Trade Review ,1894      **New MCAT Practice Tests** Kaplan, Inc,2006 Two full length practice tests make this the ideal study guide for students preparing for the MCAT Medical College Admission Test Students consistently report that taking practice tests is one of the best ways to prepare for the MCAT Perfect for use on its own or in conjunction with our comprehensive review guide Kaplan New MCAT Practice Tests Fifth Edition offers Two full length practice tests Detailed answer explanations Effective test taking strategies      **Glencoe Earth Science** Ralph M. Feather,1999 Earth science is the study of Earth and space It is the study of such things as the transfer of energy in Earth's atmosphere the evolution of landforms patterns of change that cause weather the scale and structure of stars and the interactions that occur among the water atmosphere and land Earth science in this book is divided into four specific areas of study geology meteorology astronomy and oceanography p 8 9      Heat Energy Anthea Maton,1993  
**Harcourt Science: Earth science [grade] 6, units C and D, teacher's ed** ,2000      Technometrics ,1992      **Holt Science and Technology** Holt Rinehart & Winston,2004-02      Merrill Earth Science Ralph M. Feather,Susan Leach Snyder,Dale T. Hesser,1995      **Shipman Phy SC St Gde 3/E** James T. Shipman,1979      *Teacher's Guide to Introduction to Natural Science* V. Lawrence Parsegian,1968

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