

CHAPTER 5: ATOMIC STRUCTURE

Atom consists of several subatomic particles. The main parts of an atom are:

Nucleus: central part of an atom, consists of two types of particles

- Protons:** positive charged particles. Has a relative mass of 1 atomic mass unit (u)
- Neutrons:** neutral particles. Has a relative mass of 1 atomic mass unit (u)

Electron Cloud: surrounding the nucleus, a cloud-like region divided into energy levels or electron shells.

- Electron** – negatively charged particle.

Subatomic particles – determines the properties of an atom, including its chemical behavior and reactivity

Number of protons in nucleus defines the **element's identity**.

Number of neutrons can vary within the same element, resulting in isotopes.

Number of electrons determines the atom's overall **charge**.

Number of protons = number of electrons

Isotopes

- Have different mass number but same atomic number.
- are members of a family of an element that have the same number of proton but different number of neutron.

How to Get Neutron?

$$n = A - Z$$

A = mass number

Z = atomic number

How to Get Atomic Mass?

$$A = Z + n$$

Z = atomic number

n = neutron

Ion – is an element that has different protons and electrons

Cation – positively charged ion; more protons than electrons

Anion – negatively charged ion; more electrons than protons

Electrons in Energy Levels and Sublevels

Principal Energy Level	Sublevels available	# of orbitals in sublevel	# of electrons possible in sublevel	Total electrons for energy level
(n)	(n)	(n ²)		(2n ²)
1	s	1	2	2
2	s	1	2	8
	p	3	6	
	s	1	2	
3	p	3	6	18
	d	5	10	
	s	1	2	
4	p	3	6	32
	d	5	10	
	f	7	14	

REPORTING

Relative Mass – means that is the electron has a mass of 1 unit, the proton and neutron will have a mass 2000 times that of the electron.

Proton – discovered by Eugene Goldstein

Electron – discovered by Joseph John Thompson

Neutron – discovered by Sir James Chadwick

Isotopes – atoms of the same element that have different mass (amu)

Angstrom – a common unit for describing the sizes of atoms

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Lewis Guy Rohrbaugh



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