

**CHAPTER 8 REVIEW*****Chemical Equations and Reactions*****SECTION 1: Describing Chemical reactions**

1. Match the symbol on the left with its appropriate description on the right.

        $\Delta$ 

(a) A precipitate forms.

        $\downarrow$ 

(b) A gas forms.

        $\uparrow$ 

(c) A reversible reaction occurs.

       (l)

(d) Heat is applied to the reactants.

       (aq)

(e) A chemical is dissolved in water.

        $\ddot{\cdot}$ 

(f) A chemical is in the liquid state.

2. Finish balancing the following equation:



3. In each of the following formulas, write the total number of atoms present.

       a.  $4\text{SO}_2$        b.  $8\text{O}_2$        c.  $3\text{Al}_2(\text{SO}_4)_3$ 

4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride  $\rightarrow$  aluminum fluoride + copper metal



5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



## 2neutralization Reactions Section Review Answers

**Michael J. Padilla**

## **2neutralization Reactions Section Review Answers:**

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