

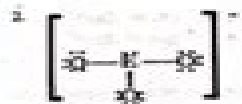
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General Certificate of Education (Adv. Level) Examination, August 2014
Chemistry - I/Two hours

- Periodic Table is Provided.
- This paper consists of 68 pages.
- Answer all the questions.
- Use of calculators is not allowed.
- Write your Index Number in the space provided in the answer sheet.
- Follow the instructions given on the back of the answer sheet carefully.
- In each of the questions 1 to 30, pick one of the alternatives form (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

Universal gas constant	$R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$
Avogadro constant	$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$
Planck's constant	$h = 6.626 \times 10^{-34} \text{ Js}$
Velocity of light	$c = 3 \times 10^8 \text{ ms}^{-1}$

1. The neutron was discovered by

- (1) Niels Bohr (2) Ernest Rutherford
 (3) James Chadwick (4) Albert Einstein
 (5) Eugen Goldstein



In the structure given above, E is an element belonging to the p-block of the Periodic Table. To which group does element E belong?

- (1) Group 13/III A (2) Group 14/IV A
 (3) Group 15/V A (4) Group 16/VI A
 (5) Group 17/VII A

3. What is the IUPAC name of the following compound?



- (1) 1-amino-6-hydroxy-2-hexen-3-one
 (2) 6-amino-1-hydroxy-4-hexen-2-one
 (3) 6-amino-2-oxo-4-hexen-1-ol
 (4) 6-hydroxy-3-oxo-2-hexenamine
 (5) 6-hydroxy-3-oxo-2-hexenylamine
4. The maximum number of electrons having quantum numbers $n = 3$ and $l = 2$ in an atom are
- (1) 2 (2) 4 (3) 6
 (4) 8 (5) 10
5. Of the following, which one has the highest boiling point?
- (1) H_2 (2) He (3) Ne
 (4) Xe (5) CH_4
6. What mass of NaCl (to the nearest gram) would contain the same total number of ions as 285 g of MgCl_2 ? ($\text{Na} = 23$, $\text{Mg} = 24$, $\text{Cl} = 35.5$)
- (1) 176 g (2) 263 g (3) 303 g
 (4) 351 g (5) 527 g
7. The solubility product of the salt XY, at 25 °C is $4.32 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-3}$. The concentration of Y in a saturated solution of XY, is
- (1) $2.0 \times 10^{-5} \text{ mol dm}^{-3}$ (2) $6.0 \times 10^{-5} \text{ mol dm}^{-3}$

- (3) $1.1 \times 10^{-5} \text{ mol dm}^{-3}$ (4) $3.8 \times 10^{-5} \text{ mol dm}^{-3}$
 (5) $4.0 \times 10^{-5} \text{ mol dm}^{-3}$

8. Which of the following reactions is not likely to take place during nitration of benzene?



9. When PCl_5 reacts with an equimolar quantity of water, the products are
- (1) POCl_3 and HCl (2) H_3PO_4 and HCl
 (3) H_2PO_4 and HCl (4) H_3PO_4 and POCl_3
 (5) POCl_3 and H_2

10. The shape and electron pair geometry of F_4ClO^- ion are respectively
- (1) trigonal bipyramidal and square pyramidal.
 (2) square pyramidal and octahedral.
 (3) trigonal bipyramidal and octahedral.
 (4) square pyramidal and trigonal bipyramidal.
 (5) octahedral and square pyramidal.

11. Which of the following statements is correct with respect to an isolated system?
- (1) Boundary of the system allows matter to exchange.
 (2) Boundary of the system does not allow matter to exchange but allows heat to exchange.
 (3) Boundary of the system allows matter or heat to exchange but it does not allow work to exchange.
 (4) Boundary of the system does not allow matter, heat and work to exchange.
 (5) Boundary of the system allows matter, heat and work to exchange.

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R Bogdan



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