

Name \_\_\_\_\_

**MULTIPLE CHOICE.** Choose the one alternative that best completes the statement or answers the question.

- The conjugate base of  $\text{HSO}_4^-$  is  
 A)  $\text{H}_2\text{SO}_4$       B)  $\text{SO}_4^{2-}$       C)  $\text{H}_3\text{SO}_4^+$       D)  $\text{HSO}_4^+$       E)  $\text{OH}^-$
- The conjugate acid of  $\text{HSO}_4^-$  is  
 A)  $\text{H}^+$       B)  $\text{H}_2\text{SO}_4$       C)  $\text{HSO}_4^+$       D)  $\text{SO}_4^{2-}$       E)  $\text{HSO}_3^+$
- The molar concentration of hydronium ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.  
 A) 7.00      B)  $1.0 \times 10^{-7}$       C) 1.00      D)  $1.0 \times 10^{-14}$       E) 0.00
- The molar concentration of hydroxide ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.  
 A) 1.00      B)  $1.0 \times 10^{-14}$       C)  $1.0 \times 10^{-7}$       D) 0.00      E) 7.00
- What is the pH of an aqueous solution at  $25^\circ\text{C}$  in which  $[\text{OH}^-]$  is 0.0025 M?  
 A) +2.60      B) -2.60      C) +11.40      D) -11.40      E) -2.25
- What is the pH of an aqueous solution at  $25^\circ$  that contains  $3.98 \times 10^{-9}$  M hydronium ion?  
 A) 7.000      B) 9.000      C) 8.400      D) 5.600      E) 3.980
- What is the concentration (in M) of hydronium ions in a solution at  $25^\circ\text{C}$  with  $\text{pH} = 4.282$ ?  
 A) 4.28      B)  $1.92 \times 10^{-10}$       C)  $1.66 \times 10^{-4}$       D) 9.71      E)  $5.22 \times 10^{-5}$
- Which solution below has the highest concentration of hydroxide ions?  
 A)  $\text{pH} = 3.21$       B)  $\text{pH} = 9.82$       C)  $\text{pH} = 7.93$       D)  $\text{pH} = 12.59$       E)  $\text{pH} = 7.00$
- What is the pH of a 0.015-M aqueous solution of barium hydroxide?  
 A) 12.18      B) 1.52      C) 12.48      D) 1.82      E) 10.35
- Sodium hydroxide is a strong base. This means that \_\_\_\_\_.  
 A) aqueous solutions of NaOH contain equal concentrations of  $\text{H}^+$  (aq) and  $\text{OH}^-$  (aq)  
 B) NaOH does not dissociate at all when it is dissolved in water  
 C) NaOH dissociates completely to  $\text{Na}^+$ (aq) and  $\text{OH}^-$ (aq) when it dissolves in water  
 D) NaOH cannot be neutralized by a weak acid  
 E) NaOH cannot be neutralized by ordinary means
- Of the following acids, \_\_\_\_\_ is not a strong acid.  
 A)  $\text{HNO}_3$       B)  $\text{HCl}$       C)  $\text{HNO}_2$       D)  $\text{HClO}_4$       E)  $\text{H}_2\text{SO}_4$
- Of the following, \_\_\_\_\_ is a weak acid.  
 A)  $\text{HBr}$       B)  $\text{HF}$       C)  $\text{HClO}_4$       D)  $\text{HCl}$       E)  $\text{HNO}_3$

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