

Name \_\_\_\_\_

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

1) The conjugate base of  $\text{HSO}_4^-$  is \_\_\_\_\_.

A)  $\text{H}_2\text{SO}_4$       B)  $\text{SO}_4^{2-}$       C)  $\text{H}_2\text{SO}_4^+$       D)  $\text{HSO}_4^+$       E)  $\text{OH}^-$

2) The conjugate acid of  $\text{HSO}_4^-$  is \_\_\_\_\_.

A)  $\text{H}^+$       B)  $\text{H}_2\text{SO}_4$       C)  $\text{HSO}_4^+$       D)  $\text{SO}_4^{2-}$       E)  $\text{HSO}_3^-$

3) The molar concentration of hydroxium ions in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.

A) 7.00      B)  $1.0 \times 10^{-7}$       C) 1.00      D)  $1.0 \times 10^{-14}$       E) 0.00

4) The molar concentration of hydroxide ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.

A) 1.00      B)  $1.0 \times 10^{-14}$       C)  $1.0 \times 10^{-7}$       D) 0.00      E) 7.00

5) What is the pH of an aqueous solution at  $25^\circ\text{C}$  in which  $[\text{OH}^-]$  is 0.0025 M?

A) +2.60      B) -2.60      C) +11.40      D) -11.40      E) -2.25

6) What is the pH of an aqueous solution at  $25^\circ\text{C}$  that contains  $3.98 \times 10^{-9}$  M hydroxium ion?

A) 7.000      B) 9.000      C) 8.400      D) 5.600      E) 3.980

7) What is the concentration (in M) of hydroxium ions in a solution at  $25^\circ\text{C}$  with pH = 4.28?

A) 4.28      B)  $1.92 \times 10^{-10}$       C)  $1.66 \times 10^4$       D) 9.71      E)  $5.22 \times 10^{-5}$

8) Which solution below has the highest concentration of hydroxide ions?

A) pH = 3.21      B) pH = 9.82      C) pH = 7.93      D) pH = 1.259      E) pH = 7.00

9) What is the pH of a 0.015-M aqueous solution of barium hydroxide?

A) 12.18      B) 1.52      C) 12.48      D) 1.82      E) 10.35

10) Sodium hydroxide is a strong base. This means that \_\_\_\_\_.

A) aqueous solutions of NaOH contain equal concentrations of  $\text{H}^+$  (aq) and  $\text{OH}^-$  (aq)  
 B) NaOH does not dissociate at all when it is dissolved in water  
 C) NaOH dissociates completely to  $\text{Na}^+$  (aq) and  $\text{OH}^-$  (aq) when it dissolves in water  
 D) NaOH cannot be neutralized by a weak acid  
 E) NaOH cannot be neutralized by ordinary means

11) Of the following acids, \_\_\_\_\_ is not a strong acid.

A)  $\text{HNO}_3$       B)  $\text{HCl}$       C)  $\text{HNO}_2$       D)  $\text{HCO}_3$       E)  $\text{H}_2\text{SO}_4$

12) Of the following, \_\_\_\_\_ is a weak acid.

A)  $\text{HBr}$       B)  $\text{HF}$       C)  $\text{HClO}_4$       D)  $\text{HCl}$       E)  $\text{HNO}_3$

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