

Advanced Placement Chemistry: 1992 Free Response Answers

- Electrode potentials are used to indicate the typical Gibbs free energy of reaction occurring in the electrode, measured in joules/mole immediately following.
- All standard redox potentials are potential within 2%.
- One point deduction for a significant figure or units error, applied only once per problem.
- Two credit earned for numerical answer without justification.

1.0 anode potential = 0.9

1.0 cathode

Initial situation is 50.00 g NaCl solution.

Electrolysis: 2.700 mole = 2 = 2.700 mole (concentration by NaCl)

$[\text{Na}^+] = 0.700 \text{ mole/liter} \rightarrow \text{pNa} = -\log 0.700 = 0.157$

$\text{pH} = -\log 0.700 \times 10^{-14} = 13.842 \text{ pH}$ and $10^{-13.842} = 10^{-13.842} \text{ M}$

= 0.157 used

Electrolysis starts by Anode Park: the actual pH standard uses 1.000 mole to 0.1000 mole H₂O₂ and then applies the 0.1000 mole to the value for a 0.1000 mole calculated. If you like the above explanation.

Electrolysis

Final situation, the ratio of NaOH/NaCl in H₂O is 2.0

$\text{pH} = 13.842 = 2.15$

$\text{pH} = 2.150 \text{ used for NaOH/NaCl concentration}$

1.000 mole x 0.1000 g/mol = 0.1000 g of NaOH/NaCl, decomposed

1.000 g x 0.1000 g = 0.1000 g of NaOH/NaCl, remaining

1.0 anode potential

$K_{\text{eq}} = 0.700 \times 10^{-14}$

= 0.700 mole x 10.00 mole = 10.00 mole²

1.0 cathode potential

The total pressure would remain at 1.000 atm. Since water added remained below 0.001 g, we can assume that there had been no temperature change, thus adding 0.001 g would not affect the equilibrium.

1.0 anode potential = 0.9

1.0 cathode

2.700 moles = 2.700 M NaCl, 27.00 mole = 2.700 x 10⁻² mole

1.0. 0.1000 M NaCl x 0.1000 mole = 0.1000 x 10⁻² mole = 0.1000 mole

0.1000 x 10⁻² M² = 1.0. 0.1000 x 10⁻² mole = 0.1000 x 10⁻² mole = 0.1000 mole

0.1000 g = 0.1000 mole = 1.000 g of NaCl

If there are decompositions of other particulate solids, the result is used to the cathode.

1.0 anode potential

$K[\text{Ca}^{2+} + 2\text{e}^{-} \rightarrow \text{Ca}]$

$[\text{Ca}] = 2\text{e}^{-} \rightarrow \text{Ca}$

$K[\text{Ca}^{2+} + 2\text{e}^{-} \rightarrow \text{Ca}] = 2\text{e}^{-}$

$\text{pH} = -\log 10^{-14}$

$\text{pH} = -\log 10^{-14}$

$\text{pH} = -\log 10^{-14}$

1.0 anode potential

$K[\text{Zn}^{2+} + 2\text{e}^{-} \rightarrow \text{Zn}]$

= 0.2 mole x 0.001000 M Zn/mole x 0.001000 mole = 0.2

= 0.2000 g = 0.2000 mole = 0.2000 g of Zn

1.0 cathode potential

$\text{pH} = -\log 10^{-14} \rightarrow -\log 10^{-14} \text{ Zn}^{2+}] + [\text{Zn}^{2+}]$

$\text{pH} = -\log 10^{-14} \text{ Zn}^{2+} + 0.2000 \text{ mole} / 0.001000 \text{ M}$

$\text{pH} = -\log 10^{-14} + 0.2000 = -0.2000$

14 Chemistry Response Answers

**Hyoung Joon Choi, Takhee Lee, Woo-
Sung Jung**

14 Chemistry Response Answers:

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