

AP[®] CHEMISTRY
2009 SCORING GUIDELINES (Form B)

Question 5 (9 points)

Answer the following questions about nitrogen, hydrogen, and ammonia.

- (a) In the boxes below, draw the complete Lewis electron-dot diagrams for N_2 and NH_3 .



The correct structures are shown in the boxes above.

Two points are earned for the correct Lewis electron-dot diagrams (1 point each).

- (b) Calculate the standard free-energy change, ΔG° , that occurs when 12.0 g of $\text{H}_2(\text{g})$ reacts with excess $\text{N}_2(\text{g})$ at 298 K according to the reaction represented below.



$$12.0 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2.0 \text{ g H}_2} \times \frac{1 \text{ mol reaction}}{3 \text{ mol H}_2} \times \frac{-34 \text{ kJ}}{1 \text{ mol reaction}} = -68 \text{ kJ}$$

One point is earned for the correct stoichiometry.

One point is earned for the correct answer.

- (c) Given that $\Delta H^\circ_{\text{rxn}}$ for the reaction is $-92.2 \text{ kJ mol}^{-1}$, which is larger, the total bond dissociation energy of the reactants or the total bond dissociation energy of the products? Explain.

$$\Delta H^\circ_{\text{rxn}} = \Sigma (\text{bond energy of the reactants}) - \Sigma (\text{bond energy of the products})$$

Based on the equation above, for $\Delta H^\circ_{\text{rxn}}$ to be negative, the total bond energy of the products must be larger than the total bond energy of the reactants.

OR

More energy is released as product bonds are formed than is absorbed as reactant bonds are broken.

One point is earned for the correct answer with the correct equation and explanation.

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Joacim Rocklöv



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