

Activity C14: Rate of a Chemical Reaction 1 (Colorimeter)

Concept	DataStudio	ScienceWorkshop (Mac)	ScienceWorkshop (Win)
Rate of reactions	C14 Reaction Rate 1.DS	C14 Reaction Rate 1	C14_REA1.SWS

Equipment Needed	Qty	Chemicals and Consumables	Qty
Colorimeter (CI-6747)	1	Hydrochloric acid, 6 molar	10 mL
Curvette (w/sensor)	1	Sodium thiosulfate, 0.2 molar	10 mL
Graduated cylinder, 10 mL	2	Water, distilled	100 mL
Protective gear	PS		

What Do You Think?

How will changing the concentrations of the reactants in a chemical reaction affect the rate of the reaction?



Take time to answer the 'What Do You Think?' question(s) in the Lab Report section.

Background

All chemical reactions occur at specific rates. The rate of a chemical reaction depends on several physical and chemical factors. These factors include:

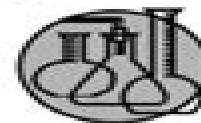
- The concentration of the reactants
- The pressure on the reaction
- The temperature of the reaction
- The presence of a catalyst

In this activity you will determine the effect of changes in concentration of the reactants on the rate of the chemical reaction. The reaction for this activity is the acidic reduction of the thiosulfate ion to sulfur and sulfur dioxide.

The equation for the reaction is:



One way to determine the effect of concentration on the rate of the reaction is to use a Colorimeter to measure the formation of the solid sulfur generated. The solid sulfur will block the light in the Colorimeter and the amount of blockage is directly proportional to the amount of sulfur in suspension.



The rate of this chemical reaction is given by the equation:

$$\text{Rate} = k [\text{thiosulfate}]^a [\text{acid}]^b$$

The letters a and b seen as exponents are numerals which can only be determined experimentally. Each reactant must be varied separately while the other is kept constant. The effect on the rate of the reaction is noted and the value of the exponent is determined in this way:

- If a change in concentration of one of the reactants has no effect, the exponent is 0.
- If doubling the concentration doubles the rate, the exponent is 1.
- If doubling the concentration quadruples the rate, the exponent is 2.

The over all order of the reaction is determined by adding a + b.

14 Chemical Reaction Unscramble

Schyrlet Cameron, Carolyn Craig



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