

2. The following questions relate to sulfur and some of its compounds.

- (a) Write the balanced equation for the combustion of  $S_s(s)$  to form  $SO_2(g)$ .
- (b) Calculate the volume of  $O_2(g)$ , measured at 1.00 atm and 298 K, that is required to completely combust a 500.0 g sample of pure  $S_s(s)$ .
- (c) A student claims that the combustion of  $S_s$  is an oxidation-reduction reaction. Justify the claim by identifying the oxidation numbers of sulfur and oxygen both before and after the reaction.
- (d) In the box below, draw a Lewis electron-dot diagram for one valid resonance structure of  $SO_2$ .

- (e) Based on the diagram you drew in part (d), what is the approximate oxygen-sulfur-oxygen bond angle in  $SO_2$ ?  
 $SO_2$  can be oxidized to form  $SO_3$ , according to the following equation.  
$$2 SO_2(g) + O_2(g) \rightarrow 2 SO_3(g) \quad \Delta H^\circ = -198 \text{ kJ/mol}_{\text{rxn}}$$
- (f) Is the value of  $\Delta S^\circ$  for the reaction represented above positive or negative? Justify your answer.  
$$2 SO_2(g) + O_2(g) \rightarrow 2 SO_3(l)$$
- (g) Is the magnitude of  $\Delta H^\circ$  for the reaction to form  $SO_3(l)$ , represented above, greater than, less than, or equal to the magnitude of  $\Delta H^\circ$  for the reaction to form  $SO_3(g)$ ? Justify your answer.
- (h) Based on the information above, how does the thermodynamic favorability of the reaction change as the temperature of the reaction system is decreased? Justify your answer.

# 14 Ap Chemistry Response Exam Questions Solution

**Mary Millhollon, Richard H. Langley**

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## **Table of Contents 14 Ap Chemistry Response Exam Questions Solution**

1. Understanding the eBook 14 Ap Chemistry Response Exam Questions Solution
  - The Rise of Digital Reading 14 Ap Chemistry Response Exam Questions Solution
  - Advantages of eBooks Over Traditional Books
2. Identifying 14 Ap Chemistry Response Exam Questions Solution
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 14 Ap Chemistry Response Exam Questions Solution
  - User-Friendly Interface
4. Exploring eBook Recommendations from 14 Ap Chemistry Response Exam Questions Solution
  - Personalized Recommendations
  - 14 Ap Chemistry Response Exam Questions Solution User Reviews and Ratings
  - 14 Ap Chemistry Response Exam Questions Solution and Bestseller Lists
5. Accessing 14 Ap Chemistry Response Exam Questions Solution Free and Paid eBooks

- 14 Ap Chemistry Response Exam Questions Solution Public Domain eBooks
- 14 Ap Chemistry Response Exam Questions Solution eBook Subscription Services
- 14 Ap Chemistry Response Exam Questions Solution Budget-Friendly Options

6. Navigating 14 Ap Chemistry Response Exam Questions Solution eBook Formats

- ePUB, PDF, MOBI, and More
- 14 Ap Chemistry Response Exam Questions Solution Compatibility with Devices
- 14 Ap Chemistry Response Exam Questions Solution Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 14 Ap Chemistry Response Exam Questions Solution
- Highlighting and Note-Taking 14 Ap Chemistry Response Exam Questions Solution
- Interactive Elements 14 Ap Chemistry Response Exam Questions Solution

8. Staying Engaged with 14 Ap Chemistry Response Exam Questions Solution

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 14 Ap Chemistry Response Exam Questions Solution

9. Balancing eBooks and Physical Books 14 Ap Chemistry Response Exam Questions Solution

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 14 Ap Chemistry Response Exam Questions Solution

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 14 Ap Chemistry Response Exam Questions Solution

- Setting Reading Goals 14 Ap Chemistry Response Exam Questions Solution
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 14 Ap Chemistry Response Exam Questions Solution

- Fact-Checking eBook Content of 14 Ap Chemistry Response Exam Questions Solution
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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