

12**STOICHIOMETRY****SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (pages 353–358)

This section explains how to calculate the amount of reactants required or product formed in a stoichiometric process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

► Using Everyday Equations (pages 353–355)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You can use the balanced equation.

2. Quantity usually means the amount of a substance expressed in grams or moles.

3. A bookcase is to be built from 3 shelves (Sh), 2 side boards (Sb), 1 top (T), 1 base (B), and 4 legs (L). Write a "balanced equation" for the construction of this bookcase.

$3\text{Sh} + 2\text{Sb} + \text{T} + \text{B} + 4\text{L} = \text{Sh}_3\text{Sb}_2\text{TB}_1\text{L}_4$.

► Using Balanced Chemical Equations (page 354)

4. Is the following sentence true or false? Stoichiometry is the calculation of quantities in chemical reactions. True

5. Calculations using balanced equations are called stoichiometric calculations.

► Interpreting Chemical Equations (pages 356–357)

6. From what elements is ammonia produced? How is it used?

Ammonia molecules are composed of nitrogen and hydrogen; it is used as a fertilizer.

7. Circle the letter of the term that tells what kind of information you CANNOT get from a chemical equation.

a. moles

d. volume

b. mass

e. number of particles

c. size of particles

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