

7.10 Compounds of sulfur :

7.10.1 Sulfur dioxide

a. Preparation :

i. **From sulfur** : Sulfur dioxide gas can be prepared by burning of sulfur in air.



ii. **From sulfite** : In the laboratory sulfur dioxide is prepared by treating aqueous solution of sodium sulfite with dilute sulfuric acid.



iii. From sulfides : (Industrial method)

Sulfur dioxide can be prepared by roasting zinc sulfide and iron pyrites.



b. Physical properties of SO_2

i. Sulfur dioxide is a colourless gas with a pungent smell.

ii. It is poisonous in nature.

iii. SO_2 is highly soluble in water and its solution in water is called sulfurous acid.

iv. It liquifies at room temperature under a pressure of 2 atm and boils at 263 K.

c. Chemical Properties :

i. **Reaction with Cl_2** : Sulfur dioxide reacts with chlorine in the presence of charcoal (catalyst) to form sulfuryl chloride.



ii. **Reaction with O_2** : Sulfur dioxide is oxidised by dioxygen in presence of vanadium (V) oxide to sulfur trioxide.



iii. **Reaction with NaOH** : Sulfur dioxide readily reacts with sodium hydroxide solution to form sodium sulfite.



iv. **Reaction with Na_2SO_3** : Sulfur dioxide reacts with sodium sulfite solution to form sodium hydrogen sulfite.



v. **Reducing property** : Sulfur dioxide acts as a reducing agent in the presence of moisture.

- Moist sulfur dioxide reduces ferric salts into ferrous salts.



- Moist sulfur dioxide decolourises acidified potassium permanganate (VII) solution.



- Moist sulfur dioxide reduces halogens to halogen acids.



d. **Structure of SO_2** : Sulfur dioxide is angular with $\text{O} = \text{S} = \text{O}$ bond angle of 119.5° .



The $\text{S} = \text{O}$ double bond arises from $\text{dx} - \text{py}$ bonding. It is a resonance hybrid of two canonical forms.

e. Uses : Sulfur dioxide is used

- In refining of petroleum and sugar.
- In bleaching wool and silk.
- As an anti-chlor, disinfectant.
- As a preservative.
- In the manufacture of H_2SO_4 , NaHSO_3 .
- Liquid SO_2 is used as a solvent to dissolve a number of organic and inorganic chemicals.

7.10.2 Sulfuric acid, H_2SO_4

a. **Preparation** : Sulfuric acid is manufactured by **Contact process**, which involves the following three steps.

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Rachel S Tattersall



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1. Understanding the eBook 12th Std Chemistry Notes
 - The Rise of Digital Reading 12th Std Chemistry Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying 12th Std Chemistry Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12th Std Chemistry Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12th Std Chemistry Notes
 - Personalized Recommendations
 - 12th Std Chemistry Notes User Reviews and Ratings
 - 12th Std Chemistry Notes and Bestseller Lists
5. Accessing 12th Std Chemistry Notes Free and Paid eBooks
 - 12th Std Chemistry Notes Public Domain eBooks
 - 12th Std Chemistry Notes eBook Subscription Services
 - 12th Std Chemistry Notes Budget-Friendly Options
6. Navigating 12th Std Chemistry Notes eBook Formats

- ePub, PDF, MOBI, and More
 - 12th Std Chemistry Notes Compatibility with Devices
 - 12th Std Chemistry Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12th Std Chemistry Notes
 - Highlighting and Note-Taking 12th Std Chemistry Notes
 - Interactive Elements 12th Std Chemistry Notes
 8. Staying Engaged with 12th Std Chemistry Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12th Std Chemistry Notes
 9. Balancing eBooks and Physical Books 12th Std Chemistry Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12th Std Chemistry Notes
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 12th Std Chemistry Notes
 - Setting Reading Goals 12th Std Chemistry Notes
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 12th Std Chemistry Notes
 - Fact-Checking eBook Content of 12th Std Chemistry Notes
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
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