

7.10 Compounds of sulfur :

7.10.1 Sulfur dioxide

a. Preparation :

i. From sulfur : Sulfur dioxide gas can be prepared by burning of sulfur in air.



ii. From sulfite : In the laboratory sulfur dioxide is prepared by treating aqueous solution of sodium sulfite with dilute sulfuric acid.



iii. From sulfides : (Industrial method)

Sulfur dioxide can be prepared by roasting zinc sulfide and iron pyrites.



b. Physical properties of SO_2 :

i. Sulfur dioxide is a colourless gas with a pungent smell.

ii. It is poisonous in nature.

iii. SO_2 is highly soluble in water and its solution in water is called sulfurous acid.

iv. It liquifies at room temperature under a pressure of 2 atm and boils at 263 K.

c. Chemical Properties :

i. Reaction with Cl_2 : Sulfur dioxide reacts with chlorine in the presence of charcoal (catalyst) to form sulfonyl chloride.



ii. Reaction with O_2 : Sulfur dioxide is oxidised by dioxygen in presence of vanadium (V) oxide to sulfur trioxide.



iii. Reaction with $NaOH$: Sulfur dioxide readily reacts with sodium hydroxide solution to form sodium sulfite.



iv. Reaction with Na_2SO_3 : Sulfur dioxide reacts with sodium sulfite solution to form sodium hydrogen sulfite.



v. Reducing property : Sulfur dioxide acts as a reducing agent in the presence of moisture.

• Moist sulfur dioxide reduces ferric salts into ferrous salts.



• Moist sulfur dioxide decolourises acidified potassium permanganate (VII) solution.



• Moist sulfur dioxide reduces halogens to halogen acids.



d. Structure of SO_2 : Sulfur dioxide is angular with $O-S-O$ bond angle of 119.5° .



The $S-O$ double bond arises from $ds-p\pi$ bonding. It is a resonance hybrid of two canonical forms.

e. Uses : Sulfur dioxide is used

- In refining of petroleum and sugar.
- In bleaching wool and silk.
- As an anti-chlor, disinfectant.
- As a preservative.
- In the manufacture of H_2SO_4 , $NaHSO_3$.
- Liquid SO_2 is used as a solvent to dissolve a number of organic and inorganic chemicals.

7.10.2 Sulfuric acid, H_2SO_4

a. Preparation : Sulfuric acid is manufactured by Contact process, which involves the following three steps.

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Rachel S Tattersall

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1. Understanding the eBook 12th Std Chemistry Notes
 - The Rise of Digital Reading 12th Std Chemistry Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying 12th Std Chemistry Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12th Std Chemistry Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12th Std Chemistry Notes
 - Personalized Recommendations
 - 12th Std Chemistry Notes User Reviews and Ratings
 - 12th Std Chemistry Notes and Bestseller Lists
5. Accessing 12th Std Chemistry Notes Free and Paid eBooks
 - 12th Std Chemistry Notes Public Domain eBooks
 - 12th Std Chemistry Notes eBook Subscription Services
 - 12th Std Chemistry Notes Budget-Friendly Options
6. Navigating 12th Std Chemistry Notes eBook Formats

- ePub, PDF, MOBI, and More
- 12th Std Chemistry Notes Compatibility with Devices
- 12th Std Chemistry Notes Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 12th Std Chemistry Notes
- Highlighting and Note-Taking 12th Std Chemistry Notes
- Interactive Elements 12th Std Chemistry Notes

8. Staying Engaged with 12th Std Chemistry Notes

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 12th Std Chemistry Notes

9. Balancing eBooks and Physical Books 12th Std Chemistry Notes

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 12th Std Chemistry Notes

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 12th Std Chemistry Notes

- Setting Reading Goals 12th Std Chemistry Notes
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 12th Std Chemistry Notes

- Fact-Checking eBook Content of 12th Std Chemistry Notes
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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