

Honors Chemistry
Extra Stoichiometry Problems

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.

a. Write and balance the chemical equation.



b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many representative particles (and what type) of silver chloride are produced?

$$\frac{39.02 \text{ g BaCl}_2}{1} \times \frac{1 \text{ mole BaCl}_2}{208.233 \text{ g BaCl}_2} \times \frac{2 \text{ moles AgCl}}{1 \text{ mole BaCl}_2} \times \frac{6.02 \times 10^{23} \text{ f. u. AgCl}}{1 \text{ mole AgCl}} = 2.256 \times 10^{23} \text{ f. u. AgCl}$$

c. If 410.8 grams of barium nitrate are produced how many grams of silver nitrate were reacted?

$$\frac{410.8 \text{ g Ba}(\text{NO}_3)_2}{1} \times \frac{1 \text{ mole Ba}(\text{NO}_3)_2}{261.335 \text{ g Ba}(\text{NO}_3)_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ mole Ba}(\text{NO}_3)_2} \times \frac{169.872 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 534.1 \text{ g AgNO}_3$$

2. Nitrogen gas is reacted with hydrogen gas to form nitrogen trihydride.

a. Write and balance the chemical equation.



b. How many liters of nitrogen trihydride are produced at STP if 80.28 grams of hydrogen gas are reacted in an excess of nitrogen?

$$\frac{80.28 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.016 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{22.4 \text{ L NH}_3}{1 \text{ mole NH}_3} = 594.7 \text{ L NH}_3$$

c. How many grams of hydrogen are needed to fully react 621.9 grams of nitrogen gas?

$$\frac{621.9 \text{ g N}_2}{1} \times \frac{1 \text{ mole N}_2}{28.014 \text{ g N}_2} \times \frac{3 \text{ moles H}_2}{1 \text{ mole N}_2} \times \frac{2.016 \text{ g H}_2}{1 \text{ mole H}_2} = 134.3 \text{ g H}_2$$

d. How many liters of nitrogen gas are needed to completely react 90.38 L of hydrogen gas at STP?

$$\frac{90.38 \text{ L H}_2}{1} \times \frac{1 \text{ mole H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mole N}_2}{3 \text{ mole H}_2} \times \frac{22.4 \text{ L N}_2}{1 \text{ mole N}_2} = 30.09 \text{ L N}_2$$

OR

$$\frac{90.38 \text{ L H}_2}{1} \times \frac{1 \text{ L N}_2}{3 \text{ L H}_2} = 30.09 \text{ L N}_2$$

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Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change. *Introduction to Chemistry, Laboratory Manual* T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities. These themes involve and engage the students in their own learning processes they are challenged to be active. The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics. The reorganization presents many core concepts in the first five chapters preparing students for later chapters. In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students.

Stoichiometry Unit Project Luann Marie Decker, 1998 **Chemistry** John McMurry, Robert C. Fay, 2001 A colorful pedagogically enhanced standard textbook for the introductory course. It begins with atomic structure proceeds next to bonding and molecules then to bulk physical properties of substances and ends with a study of chemical properties. Each chapter concludes with a brief description of an interesting application or extension of the chapter subject a summary a list of key words and a large number of problems. Many student oriented supplements are available. Annotation copyright by Book News Inc Portland OR.

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