

Physics

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(Chapter 11) (Thermal Properties of Matter) (Exercises) (Class 11)

Question 11.15:

Given below are observations on molar specific heats at room temperature of some common gases.

Gas	Molar specific heat (C_v) (cal mol $^{-1}$ K $^{-1}$)
Hydrogen	4.87
Nitrogen	4.97
Oxygen	5.02
Nitric oxide	4.99
Carbon monoxide	5.01
Chlorine	6.17

The measured molar specific heats of these gases are markedly different from those for monatomic gases. Typically, molar specific heat of a monatomic gas is 2.92 cal/mol K. Explain this difference. What can you infer from the somewhat larger (than the rest) value for chlorine?

Answer 11.15:

The gases listed in the given table are diatomic. Besides the translational degree of freedom, they have other degrees of freedom (modes of motion).

Heat must be supplied to increase the temperature of these gases. This increases the average energy of all the modes of motion. Hence, the molar specific heat of diatomic gases is more than that of monatomic gases.

If only rotational mode of motion is considered, then the molar specific heat of a diatomic gas

$$=\frac{5}{2}R \Rightarrow \frac{5}{2} \times 1.98 = 4.95 \text{ cal mol}^{-1} \text{ K}^{-1}$$

With the exception of chlorine, all the observations in the given table agree with $\left(\frac{5}{2}R\right)$.

This is because at room temperature, chlorine also has vibrational modes of motion besides rotational and translational modes of motion.

Question 11.16:

A child running a temperature of 101°F is given an antipyrrin (i.e. a medicine that lowers fever) which causes an increase in the rate of evaporation of sweat from his body. If the fever is brought down to 98 °F in 20 minutes, what is the average rate of extra evaporation caused, by the drug. Assume the evaporation mechanism to be the only way by which heat is lost. The mass of the child is 30 kg. The specific heat of human body is approximately the same as that of water, and latent heat of evaporation of water at that temperature is about 580 cal g $^{-1}$.

Answer 11.16:

Initial temperature of the body of the child, $T_1 = 101^\circ\text{F}$

Final temperature of the body of the child, $T_2 = 98^\circ\text{F}$

$$\text{Change in temperature, } \Delta T = \left[(101 - 98) \times \frac{5}{9} \right]^\circ\text{C}$$

Time taken to reduce the temperature, $t = 20 \text{ min}$

Mass of the child, $m = 30 \text{ kg} = 30 \times 10^3 \text{ g}$

Specific heat of the human body = Specific heat of water = $c = 1000 \text{ cal/kg/}^\circ\text{C}$

Latent heat of evaporation of water, $L = 580 \text{ cal g}^{-1}$

$$\text{The heat lost by the child is given as: } \Delta Q = mc\Delta T = 30 \times 1000 \times (101 - 98) \times \frac{5}{9} = 50000 \text{ cal}$$

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1. Understanding the eBook 11th Physics All Lesson Target
 - The Rise of Digital Reading 11th Physics All Lesson Target
 - Advantages of eBooks Over Traditional Books
2. Identifying 11th Physics All Lesson Target
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 11th Physics All Lesson Target
 - User-Friendly Interface
4. Exploring eBook Recommendations from 11th Physics All Lesson Target

- Personalized Recommendations
- 11th Physics All Lesson Target User Reviews and Ratings
- 11th Physics All Lesson Target and Bestseller Lists

5. Accessing 11th Physics All Lesson Target Free and Paid eBooks
 - 11th Physics All Lesson Target Public Domain eBooks
 - 11th Physics All Lesson Target eBook Subscription Services
 - 11th Physics All Lesson Target Budget-Friendly Options
6. Navigating 11th Physics All Lesson Target eBook Formats
 - ePub, PDF, MOBI, and More
 - 11th Physics All Lesson Target Compatibility with Devices
 - 11th Physics All Lesson Target Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 11th Physics All Lesson Target
 - Highlighting and Note-Taking 11th Physics All Lesson Target
 - Interactive Elements 11th Physics All Lesson Target
8. Staying Engaged with 11th Physics All Lesson Target
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 11th Physics All Lesson Target
9. Balancing eBooks and Physical Books 11th Physics All Lesson Target
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 11th Physics All Lesson Target
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 11th Physics All Lesson Target
 - Setting Reading Goals 11th Physics All Lesson Target
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 11th Physics All Lesson Target

- Fact-Checking eBook Content of 11th Physics All Lesson Target
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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