

Physics

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(Chapter 12) (Kinetic Theory) (Exercises) (Class 11)

Question 12.6:

Estimate the total number of air molecules (inclusive of oxygen, nitrogen, water vapour and other constituents) in a room of capacity 25.0 m^3 at a temperature of 27°C and 1 atm pressure.

Answer 12.6:

Volume of the room, $V = 25.0 \text{ m}^3$

Temperature of the room, $T = 27^\circ\text{C} = 300 \text{ K}$

Pressure in the room, $P = 1 \text{ atm} = 1 \times 1.013 \times 10^5 \text{ Pa}$

The ideal gas equation relating pressure (P), Volume (V), and absolute temperature (T) can be written as: $PV = k_B NT$

Where, k_B is Boltzmann constant $= 1.38 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$

N is the number of air molecules in the room.

$$\therefore N = \frac{PV}{k_B T} = \frac{1.013 \times 10^5 \times 25}{1.38 \times 10^{-23} \times 300} = 6.11 \times 10^{26}$$

Therefore, the total number of air molecules in the given room is 6.11×10^{26} .

Question 12.7:

Estimate the average thermal energy of a helium atom at (i) room temperature (27°C), (ii) the temperature on the surface of the Sun (6000 K), (iii) the temperature of 10 million kelvin (the typical core temperature in the case of a star).

Answer 12.7:

At room temperature, $T = 27^\circ\text{C} = 300 \text{ K}$

k is Boltzmann constant $= 1.38 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$

$$\text{Average thermal energy} = \frac{3}{2} kT = \frac{3}{2} \times 1.38 \times 10^{-23} \times 300 = 6.21 \times 10^{-21} \text{ J}$$

Hence, the average thermal energy of a helium atom at room temperature (27°C) is $6.21 \times 10^{-21} \text{ J}$.

On the surface of the sun, $T = 6000 \text{ K}$

$$\text{Average thermal energy} = \frac{3}{2} kT = \frac{3}{2} \times 1.38 \times 10^{-23} \times 600 = 1.241 \times 10^{-19} \text{ J}$$

Hence, the average thermal energy of a helium atom on the surface of the sun is $1.241 \times 10^{-19} \text{ J}$.

At temperature, $T = 10^7 \text{ K}$

$$\text{Average thermal energy} = \frac{3}{2} kT = \frac{3}{2} \times 1.38 \times 10^{-23} \times 10^7 = 2.07 \times 10^{-16} \text{ J}$$

Hence, the average thermal energy of a helium atom at the core of a star is $2.07 \times 10^{-16} \text{ J}$.

Question 12.8:

Three vessels of equal capacity have gases at the same temperature and pressure. The first vessel contains neon (monatomic), the second contains chlorine (diatomic), and the third contains uranium hexafluoride (polyatomic). Do the vessels contain equal number of respective molecules? Is the root mean square speed of molecules the same in the three cases? If not, in which case is v_{rms} the largest?

Answer 12.8:

Yes. All contain the same number of the respective molecules.

No. The root mean square speed of neon is the largest.

Since the three vessels have the same capacity, they have the same volume.

Hence, each gas has the same pressure, volume, and temperature.

According to Avogadro's law, the three vessels will contain an equal number of the respective molecules. This number is equal to Avogadro's number, $N = 6.023 \times 10^{23}$.

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