

Science

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(Chapter – 1) (Chemical Reactions and Equations) (Exercises) (Class 10)

Question 1:

Which of the statements about the reaction below are incorrect?



- (a) Lead is getting reduced.
- (b) Carbon dioxide is getting oxidised.
- (c) Carbon is getting oxidised.
- (d) Lead oxide is getting reduced.

(i) (a) and (b)

(ii) (a) and (c)

(iii) (a), (b) and (c)

(iv) all

Answer 1:

(i) (a) and (b)

Question 2:

The above reaction is an example of a

- (a) combination reaction.
- (c) decomposition reaction.



- (b) double displacement reaction.
- (d) displacement reaction.

Answer 2:

(d) displacement reaction.

Question 3:

What happens when dilute hydrochloric acid is added to iron filings? Tick the correct answer.

- (a) Hydrogen gas and iron chloride are produced.
- (c) No reaction takes place.

- (b) Chlorine gas and iron hydroxide are produced.
- (d) Iron salt and water are produced.

Answer 3:

(a) Hydrogen gas and iron chloride are produced.

Question 4:

What is a balanced chemical equation? Why should chemical equations be balanced?

Answer 4:

A reaction which has an equal number of atoms of all the elements on both sides of the chemical equation is called a balanced chemical equation.

The law of conservation of mass states that mass can neither be created nor destroyed. Hence, in a chemical reaction, the total mass of reactants should be equal to the total mass of the products. It means that the total number of atoms of each element should be equal on both sides of a chemical equation. Hence, it is for this reason that chemical equations should be balanced.

Question 5:

Translate the following statements into chemical equations and then balance them.

- (a) Hydrogen gas combines with nitrogen to form ammonia.
- (b) Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
- (c) Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
- (d) Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.

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 - The Rise of Digital Reading 10th Cbse Science Guide Of Physics
 - Advantages of eBooks Over Traditional Books
2. Identifying 10th Cbse Science Guide Of Physics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 10th Cbse Science Guide Of Physics
 - User-Friendly Interface
4. Exploring eBook Recommendations from 10th Cbse Science Guide Of Physics
 - Personalized Recommendations
 - 10th Cbse Science Guide Of Physics User Reviews and Ratings
 - 10th Cbse Science Guide Of Physics and Bestseller Lists

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 - Highlighting and Note-Taking 10th Cbse Science Guide Of Physics
 - Interactive Elements 10th Cbse Science Guide Of Physics
8. Staying Engaged with 10th Cbse Science Guide Of Physics
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 10th Cbse Science Guide Of Physics
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 10th Cbse Science Guide Of Physics
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals 10th Cbse Science Guide Of Physics
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 10th Cbse Science Guide Of Physics
 - Fact-Checking eBook Content of 10th Cbse Science Guide Of Physics
 - Distinguishing Credible Sources
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