

CHEMISTRY 103 – Practice Sheet #4 – Naming Compounds and Writing Formulas

Chapter 2

Prepared by Dr. Tony Jacob

<http://www.chem3tsc.edu/nastaiclc> (Resource page)

1. Name the compounds below. Includes molecular and ionic compounds, polyatomic ions, acids, and transition metals.

a. Na_2S sodium sulfideb. CaO calcium oxidec. N_2O dinitrogen monoxided. K_3N potassium nitridee. NO nitrogen oxidef. Li_3P lithium phosphideg. NaNO_3 sodium nitrateh. $\text{Ca}_3(\text{PO}_4)_2$ calcium phosphatei. MgF_2 magnesium fluoridej. $(\text{NH}_4)_2\text{SO}_4$ ammonium sulfatek. P_2O_9 diphosphorus pentoxidel. NaOH sodium hydroxidem. HBr hydrogen bromiden. $\text{Sc}(\text{OH})_3$ scandium III hydroxideo. HNO_3 nitric acidp. NH_4NO_3 ammonium nitrateq. FeCl_3 ferric chlorider. $(\text{NH}_4)_2\text{O}$ ammonium oxides. CuO copper oxidet. HCl hydrochloric acidu. NaHCO_3 sodium bicarbonatev. $\text{Al}(\text{C}_2\text{H}_5\text{O}_2)_3$ aluminum acetatew. H_2SO_4 dihydrogen sulfatex. HI hydroiodic acidy. SO_3 sulfur trioxidez. $\text{HC}_2\text{H}_3\text{O}_2$ acetic acidaa. $\text{Ca}(\text{OH})_2$ calcium hydroxidebb. $\text{Mn}(\text{ClO}_4)_3$ manganese III perchloratecc. HClO_4 perchlorate aciddd. $\text{Zn}(\text{CN})_2$ zinc II cyanideee. $\text{NaC}_2\text{H}_3\text{O}_2$ sodium acetateff. K_2SO_4 potassium sulfategg. H_2CO_3 dihydrogen carbonatehh. H_3PO_4 trihydrogen phosphateii. HF hydrofluoric acidjj. CH_3COOH acetic acidkk. LiOH lithium hydroxidell. H_3O^+ trihydroxide ionmm. NH_3 ammoniann. CH_4 methane

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