

## AP BIO: Unit One Study Guide

- Organic compounds: contain covalently bonded carbon; examples include lipids, proteins, nucleic acids and carbohydrates
- Functional groups: amino, carbenyl, carboxyl, hydroxyl, phosphate, and sulfhydryl (*recognize structure*)
- Lipids: non-polymer molecules used for long term energy storage (especially in animals), water-proofing, and insulation
  - o Fat/Oil: glycerol + 3 fatty acids; capital E shape
    - Saturated fat: bad for you; animals and some plants have it; solidifies at room temperature
    - Unsaturated fat: better for you, plants have it; liquefies at room temperature
  - o Steroids: lipids whose structures resemble chicken-wire fence; include cholesterol and sex hormones
  - o Phospholipids: glycerol + 2 fatty acids + 1 phosphate group; make up all membrane bilayers in cells; have hydrophobic interiors (the tails) and hydrophilic exteriors (the heads)
- Carbohydrates: used by cells for energy and structure; monomers are monosaccharides (glucose) or single sugars, disaccharides have two sugars (sucrose, maltose, lactose), and starches are long polymer chains of sugar; storage polysaccharides (starch [plants], glycogen [animals]), structural polysaccharides (chitin [fungal cell walls], cellulose [plant cell walls])
- Proteins: made of amino acid monomers; amino acids consist of an R group attached to an amino and a carboxyl; the amino group of one amino acid can be bonded to the carboxyl group of another amino acid through dehydration synthesis to create a peptide bond and build the protein polymer; proteins serve many functions (e.g., transport, enzymes, cell signals, receptor molecules, structural components, and channels)
  - o Structure: Primary – amino acid polymer made at the ribosome; Secondary – beta pleats or alpha helices; Tertiary – interactions between the R groups of the polymer; Quaternary – interactions with other separate polymer chains (find a buddy)
  - o Enzymes: catalytic proteins that react in an induced-fit fashion with substrates to speed up the rate of reactions by lowering the activation energy; effectiveness is affected by changes in pH, temperature, salinity, etc due to denaturing of structure; examples – pepsin digests proteins and amylase digests starch within the digestive tract
- Nucleic Acids: DNA and RNA; carry the genetic code; monomers are nucleotides consisting of a phosphate group, 5-carbon sugar, and one of five nitrogen bases (adenine, guanine, cytosine, thymine, and uracil)
- ATP: Adenosine triphosphate molecule that acts as a temporary energy storage molecule (battery) for cellular reactions and processes; consisting of an adenosine molecule covalently bonded to three phosphate groups, the addition of the third phosphate group through dehydration synthesis stores small amounts of energy while the removal of the third phosphate group through hydrolysis releases small amounts of energy
- pH: logarithmic scale <7 acidic, 7 neutral, >7 basic (alkaline); pH 4 is 10 times more acidic than pH5; buffers are enzymes that neutralize pH
- Reaction types
  - o Hydrolysis: breaks down polymers by adding water
  - o Dehydration synthesis: creates polymers by removing water
  - o Endergonic reaction: reaction that requires energy input (feels cold); makes bonds
  - o Exergonic reaction: reaction that gives off energy (feels hot); breaks bonds
- Bond types
  - o Covalent bond: sharing electrons; strong
  - o Ionic bond: transferring electrons; strong
  - o Hydrogen bond: temporary attraction between polar molecules; weak; always represented by a dotted line

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**Mark Anestis, Kellie Cox**



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