

# Gravimetric Analysis of Chloride in Solution

Experiment 10  
Analytical Chemistry  
Lab 10

## Introduction

This experiment is designed to determine the concentration of chloride ions in a solution using gravimetric analysis. The method involves the precipitation of chloride ions as silver chloride ( $\text{AgCl}$ ) by adding a known amount of silver nitrate ( $\text{AgNO}_3$ ) to the solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is designed to be performed in a laboratory setting and is suitable for students with a basic understanding of chemistry. The experiment is designed to be performed in a laboratory setting and is suitable for students with a basic understanding of chemistry. The experiment is designed to be performed in a laboratory setting and is suitable for students with a basic understanding of chemistry.

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## Procedure

The procedure involves the precipitation of chloride ions as silver chloride ( $\text{AgCl}$ ) by adding a known amount of silver nitrate ( $\text{AgNO}_3$ ) to the solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

## Data Analysis

Sample #1	Sample #2
Mass of sample	0.25g
Mass of precipitate	0.45g
Mass of chloride ion	0.20g

The mass of chloride ion in the sample is 0.20g.

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# Analysis Of Chloride Lab Report

**Shasha Hu**



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