

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the acids and bases that are indicated in the spaces provided. Be sure to include the correct state symbols.

1. What substance is _____, a weak hydrogen ion concentration of $1.0 \times 10^{-4} \text{ M}$ is the concentration of this acidic substance in water and equal to _____ M .
2. Hydrochloric acid ionizes from _____ to H^+ and _____.
3. _____, a _____, ionizes from _____ to _____ and _____.
4. _____, a _____, ionizes from _____ to _____ and _____.
5. _____, a _____, ionizes from _____ to _____ and _____.
6. _____, a _____, ionizes from _____ to _____ and _____.
7. _____, a _____, ionizes from _____ to _____ and _____.

8. _____
9. _____
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16. _____
17. _____
18. _____

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Christian G. Meyer



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