

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                       |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution      |
| 4. autoionization | 5. acid rain                | 6. acid-base reaction |
|                   | 7. acid solution            | 8. pH                 |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the values indicated that are indicated in the space provided. Be sure to include units when appropriate.

1. What substance is  $\text{H}^+$  in water? Hydrogen ion.
2. What is the pH of a 0.1 M solution of HCl? 1.0.
3. What is the pH of a 0.1 M solution of NaOH? 13.0.
4. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.
5. What is the pH of a 0.1 M solution of  $\text{Na}_2\text{SO}_4$ ? 7.0.
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28. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.
29. What is the pH of a 0.1 M solution of  $\text{Na}_2\text{SO}_4$ ? 7.0.
30. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.
31. What is the pH of a 0.1 M solution of  $\text{Na}_2\text{SO}_4$ ? 7.0.
32. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.
33. What is the pH of a 0.1 M solution of  $\text{Na}_2\text{SO}_4$ ? 7.0.
34. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.
35. What is the pH of a 0.1 M solution of  $\text{Na}_2\text{SO}_4$ ? 7.0.
36. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.
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100. What is the pH of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? 1.0.

1.  $\text{H}^+$  in water

2. 1.0

3. 13.0

4. 1.0

5. 7.0

6. 1.0

7. 7.0

8. 1.0

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## **Table of Contents 19 2 Hydrogen Ions And Acidity Answer Key**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Answer Key
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Answer Key
  - Personalized Recommendations
  - 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings

- 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists
- 5. Accessing 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks
  - 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
  - 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
  - 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
  - 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Answer Key
  - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Answer Key
  - Interactive Elements 19 2 Hydrogen Ions And Acidity Answer Key
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Answer Key
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Answer Key
  - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Answer Key
  - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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