

Experiment 11**Data and Calculations: The Atomic Spectrum of Hydrogen****A. Calculation of the Energy Levels of the Hydrogen Atom**

Energies are to be calculated from Equation 6 for the 10 lowest energy states.

Table 11.2

Quantum number, n	Energy, E_n [kJ/mol]	Quantum number, n	Energy, E_n [kJ/mol]
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

B. Calculation of the Wavelengths of the Lines in the Hydrogen Atomic Spectrum

In the upper half of each box write ΔE , the difference in energy in kJ/mole between E_{n_u} and E_{n_l} . In the lower half of the box, write the λ in nm associated with that value of ΔE .

Table 11.3

n_u	6	5	4	3	2	1
n_l						
1	1275.6 kJ/mol 93.78 nm	1254.50 kJ/mol 94.97 nm	1230.04 kJ/mol 97.25 nm	1166.25 kJ/mol 102.57 nm	984.03 kJ/mol 121.50 nm	
2	291.57 kJ/mol 410.23 nm	275.53 kJ/mol 434.12 nm	246.01 kJ/mol 486.10 nm	182.23 kJ/mol 556.46 nm		
3	109.34 kJ/mol 1094.87 nm	93.5 kJ/mol 1251.17 nm	63.78 kJ/mol 1515.61 nm			
4	45.50 kJ/mol 2675.70 nm	29.52 kJ/mol 4052.48 nm				
5	16.04 kJ/mol 7458.04 nm					

$$\Delta E = E_{n_u} - E_{n_l}$$

$$\lambda \text{ [nm]} = \frac{1.19627 \times 10^5}{\Delta E}$$

(continued on following page)

Atomic Spectrum Of Hydrogen Lab Answer Key

Léo Pomar



Atomic Spectrum Of Hydrogen Lab Answer Key:

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Atomic Spectrum Of Hydrogen Lab Answer Key**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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